Advance Study Assignment: Resolution of Matter Into Pure Substances. II. Fractional Crystalization

- L. Using Figure 3.1, determine
 - the number of grams of CuSO₄ · 5H₂O that will dissolve in 100 g of H₂O at 100°C.

180 _1000, SHO

the number of grams of water required to dissolve 4.0 g of Caso, - SH.O at 100°C. (Hisse your answer to Problem 1(a) gives you the needed conversion factor for g Cuso, - SH.O at 100°C, Grams of water = (g) ren mass of Cuso 4 · 5H2O)

At 100°C, Grams of water (40 Cuso 4 · 5H3O) * (1000g HaO) / 1809 Cuso 4 · 5H3O) * (1000g HaO) / 1809 Cuso 4 · 5H3O)

At 100°C, Grams of water (40 Cuso 4 · 5H3O) * (1000g HaO) / 1809 Cuso 4 · 5H3O)

At 100°C, Grams of water (40 Cuso 4 · 5H3O) * (1000g HaO) / 1809 Cuso 4 · 5H3O)

c. the number of grams of water required to dissolve 18 g KNO₂ at 100°C.

From the given figure, at 100°C, 2 to g of knos is soluble in bog water. To dissolve 189 knos.

(150g Kno3) . (100g H20/2HO Kno3)
at 100°C grams of water=2-50g

7 50 EHO

4. Consider of grows of water required to dissolve a mixture containing 18 g KNO₃ and 4.0 g CuSO₄ = 5.5 ft, assembling that the solutions of one substance is not affected by the presence of another.

At 1000, 2-29 of water is required to dissolve 409 Cuso-54,0 and 759 of water is required to dissolve 18.09 1403

Grams of water required to dissolve 409 of cuso-5420 and 1809 of 4003 at 100°C is 7.59

Resolution Of Matter Into Pure Substances Answers

RD Boyd

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Chemical Principles in the Laboratory Emil J. Slowinski,1996 Provides a series of experiments designed to teach students the available experimental methods the proper design of experiments and the interpretation of experimental results

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