

## CHAPTER 15 REVIEW

***Acid-Base Titration and pH*****SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. Calculate the following values without using a calculator.

- \_\_\_\_\_ a. The  $[\text{H}_3\text{O}^+]$  is  $1 \times 10^{-6}$  M in a solution. Calculate the  $[\text{OH}^-]$ .
- \_\_\_\_\_ b. The  $[\text{H}_3\text{O}^+]$  is  $1 \times 10^{-9}$  M in a solution. Calculate the  $[\text{OH}^-]$ .
- \_\_\_\_\_ c. The  $[\text{OH}^-]$  is  $1 \times 10^{-12}$  M in a solution. Calculate the  $[\text{H}_3\text{O}^+]$ .
- \_\_\_\_\_ d. The  $[\text{OH}^-]$  in part c is reduced by half, to  $0.5 \times 10^{-12}$  M. Calculate the  $[\text{H}_3\text{O}^+]$ .
- \_\_\_\_\_ e. The  $[\text{H}_3\text{O}^+]$  and  $[\text{OH}^-]$  are \_\_\_\_\_ (directly, inversely, or not) proportional in any system involving water.

2. Calculate the following values without using a calculator.

- \_\_\_\_\_ a. The pH of a solution is 2.0. Calculate the pOH.
- \_\_\_\_\_ b. The pOH of a solution is 4.73. Calculate the pH.
- \_\_\_\_\_ c. The  $[\text{H}_3\text{O}^+]$  in a solution is  $1 \times 10^{-3}$  M. Calculate the pH.
- \_\_\_\_\_ d. The pOH of a solution is 5.0. Calculate the  $[\text{OH}^-]$ .
- \_\_\_\_\_ e. The pH of a solution is 1.0. Calculate the  $[\text{OH}^-]$ .

3. Calculate the following values.

- \_\_\_\_\_ a. The  $[\text{H}_3\text{O}^+]$  is  $2.34 \times 10^{-5}$  M in a solution. Calculate the pH.
- \_\_\_\_\_ b. The pOH of a solution is 3.5. Calculate the  $[\text{OH}^-]$ .
- \_\_\_\_\_ c. The  $[\text{H}_3\text{O}^+]$  is  $4.6 \times 10^{-8}$  M in a solution. Calculate the  $[\text{OH}^-]$ .

**PROBLEMS** Write the answer on the line to the left. Show all your work in the space provided.4.  $[\text{H}_3\text{O}^+]$  in an aqueous solution =  $2.3 \times 10^{-3}$  M.

- \_\_\_\_\_ a. Calculate  $[\text{OH}^-]$  in this solution.

# Review Acid Base Titration And Ph Answer

**Anthony S. Fauci**



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