



Name: _____

Date: _____

Student Exploration: Ionic Bonds

Vocabulary: chemical family, electron affinity, ion, ionic bond, metal, nonmetal, octet rule, shell, valence electron

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. Nate and Clara are drawing pictures with markers. There are 8 markers in a set. Nate has 9 markers and Clara has 7. What can Nate and Clara do so that each of them has a full set?

Nate gives one marker to Clara.

2. Maggie is sitting at a table with Fred and Florence. Maggie has 10 markers, but Fred and Florence each have only 7 markers. How can they share markers so each has 8?

Maggie gives one to Fred and then Florence.

Gizmo Warm-up

Just like students sharing markers, atoms sometimes share or swap electrons. By doing this, atoms form bonds. The *Ionic Bonds* Gizmo allows you to explore how **ionic bonds** form.

To begin, check that **Sodium (Na)** and **Chlorine (Cl)** are selected from the menus at right. Click **Play** (▶) to see electrons orbiting the nucleus of each atom. (Note: These atom models are simplified and not meant to be realistic.)



1. Each atom consists of a central nucleus and several **shells** that contain electrons. The outermost electrons are called **valence electrons**. (Inner electrons are not shown.)

How many valence electrons does each atom have? Sodium: **1** Chlorine: **7**

2. Click **Pause** (⏸). Elements can be classified as **metals** and **nonmetals**. Metals do not hold on to their valence electrons very tightly, while nonmetals hold their electrons tightly. **Electron affinity** is a measure of how tightly the valence electrons are held.

- A. Try pulling an electron away from each atom. Based on this experiment, which atom is a metal? **Sodium** Which is a nonmetal?

Chlorine



Review And Reinforce Ionic Bonds Answer Key

Michael J. Padilla



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Prentice Hall Science Explorer Michael J. Padilla, 2002

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1. Understanding the eBook Review And Reinforce Ionic Bonds Answer Key
 - The Rise of Digital Reading Review And Reinforce Ionic Bonds Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Review And Reinforce Ionic Bonds Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review And Reinforce Ionic Bonds Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review And Reinforce Ionic Bonds Answer Key
 - Personalized Recommendations

- Review And Reinforce Ionic Bonds Answer Key User Reviews and Ratings
- Review And Reinforce Ionic Bonds Answer Key and Bestseller Lists
- 5. Accessing Review And Reinforce Ionic Bonds Answer Key Free and Paid eBooks
 - Review And Reinforce Ionic Bonds Answer Key Public Domain eBooks
 - Review And Reinforce Ionic Bonds Answer Key eBook Subscription Services
 - Review And Reinforce Ionic Bonds Answer Key Budget-Friendly Options
- 6. Navigating Review And Reinforce Ionic Bonds Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Review And Reinforce Ionic Bonds Answer Key Compatibility with Devices
 - Review And Reinforce Ionic Bonds Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review And Reinforce Ionic Bonds Answer Key
 - Highlighting and Note-Taking Review And Reinforce Ionic Bonds Answer Key
 - Interactive Elements Review And Reinforce Ionic Bonds Answer Key
- 8. Staying Engaged with Review And Reinforce Ionic Bonds Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review And Reinforce Ionic Bonds Answer Key
- 9. Balancing eBooks and Physical Books Review And Reinforce Ionic Bonds Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review And Reinforce Ionic Bonds Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Review And Reinforce Ionic Bonds Answer Key
 - Setting Reading Goals Review And Reinforce Ionic Bonds Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review And Reinforce Ionic Bonds Answer Key
 - Fact-Checking eBook Content of Review And Reinforce Ionic Bonds Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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