

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

\_\_\_\_\_ **lattice energy** \_\_\_\_\_

a. ionic bond

\_\_\_\_\_ **bond energy** \_\_\_\_\_

b. covalent bond

\_\_\_\_\_ **enthalpy of vaporization** \_\_\_\_\_

c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

\_\_\_\_\_ **ionic** \_\_\_\_\_

a. H—F

\_\_\_\_\_ **nonpolar covalent** \_\_\_\_\_

d. H—H

\_\_\_\_\_ **ionic** \_\_\_\_\_

b. Na—Cl

\_\_\_\_\_ **polar covalent** \_\_\_\_\_

e. H—C

\_\_\_\_\_ **polar covalent** \_\_\_\_\_

c. H—O

\_\_\_\_\_ **polar covalent** \_\_\_\_\_

f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen  
atom and the unshared electron pair of a strongly electronegative atom such as O,  
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,  
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally  
intermolecular, while ionic and covalent bonds occur between ions or atoms  
respectively.

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.  
Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making  
water a liquid at room temperature. Hydrogen bonding exists between water  
molecules, but not between hydrogen sulfide molecules.

# Review Chemical Bonding Section 5 Answers

**Lei Shi**



## Review Chemical Bonding Section 5 Answers:

**A Complete Preparation for the MCAT** Aftab S. Hassan, James L. Flowers, 1992 This guide for MCAT preparation applies the principles of active and problem based learning to an updated review of content and skills with models for enhanced problem solving and critical thinking abilities There are details on setting up a self managed study programme with guidelines for time management and stress management All areas tested on the exam are covered verbal reasoning physical science writing sample biological sciences with practice questions to chart progress *Science Tutor: Chemistry, Grades 7 - 8* Raham, 2008-08-28 Connect students in grades 7 and up with science using Science Tutor Chemistry This effective 48 page resource provides additional concept reinforcement for students who struggle in chemistry Each lesson in this book contains an Absorb section to instruct and simplify concepts and an Apply section to help students grasp concepts on their own The book covers topics such as matter physical and chemical changes mixtures and solutions the periodic table atomic structure and radioactivity It is great for use in the classroom and at home *Science Tutor: Chemistry, Grades 7 - 12* Gary Raham, 2005-01-03 Introduces new chemistry concepts and provides activities so that students can practice and grasp the concepts Key terms are highlighted in the text as well as in a comprehensive glossary Answer keys are included *Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter **Basic Concepts of Chemistry, Study Guide** Leo J. Malone, 2000 The text's three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems *Study Guide to accompany Basic Concepts of Chemistry,*

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