

Revision - Physical Chemistry - 3a) Energetics

Define the following reactions and give an example of each

Endothermic absorbs heat to the environment e.g. Boiling

Exothermic Takes in heat e.g. Melting ice

Explain what experiment you can do to determine how much energy is given out during neutralization

Put 100cm³ in a polystyrene cup & measure temperature at the start & end using a thermometer

Draw a labelled energy level diagram to show an exothermic reaction.



1. Excess iron powder was added to 100.0 cm³ of 0.200 mol dm⁻³ copper(II) sulfate solution in a calorimeter. The reaction equation was as follows. $\text{Fe (s)} + \text{CuSO}_4\text{(aq)} \rightarrow \text{FeSO}_4\text{(aq)} + \text{Cu (s)}$
The maximum temperature rise was 7.5 °C. Determine the heat energy change of the reaction, in kJ.

$$Q = m \times c \times \Delta T$$

$$Q = 100\text{g} \times 4.18 \text{ J/g}^\circ\text{C} \times 7.5^\circ\text{C} = -3.13 \text{ kJ}$$

2. 1.023 g of propan-1-ol ($M_r = 60.11 \text{ g mol}^{-1}$) was burned in a spirit burner and used to heat 200 g of water in a copper calorimeter. The temperature of the water rose by 30 °C. Calculate the heat energy change for the combustion of propan-1-ol using this data.

$$Q = 200\text{g} \times 4.18 \times 30^\circ\text{C} = -25080 \text{ J} \approx -25 \text{ kJ}$$

3. A chemist uses a Bunsen burner to heat some ethanol in a beaker. If the energy input is 2.5 kJ and the temperature change is 21 °C to 32 °C, find the mass of ethanol that was heated in kg.

$$m \Delta T = \frac{2500}{2400 \times 11} = 0.095 \text{ kg / 95g}$$

3. A spinning bicycle wheel needs to be brought to a stop by a rubber brake block. Assuming the wheel has 500 J of kinetic energy that is all absorbed by the brake and that the mass of the block is 25g, how much will the temperature of the block rise by in °C?

$$\Delta T = \frac{500}{0.025 \times 2000} = 10^\circ\text{C}$$

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