## Review Sheet - Chemistry, Level 3 - Ch. 14: Liquids & Solids

Name:	ANSWERS	Date:	Period:
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This test covers materials in the textbook in Chapters 14 (pp. 456-491). You are always responsible for the topics on the previous tests.

#### Test Expectations

The student should be able:

- Describe how the Kinetic-Molecular Theory of Matter predicts the properties of substances
- Analyze the forces and relative strengths of the forces between particles in a solid, liquid or gas
  and compare the substances in terms of basic properties such as melting point, boiling point,
  viscosity, surface tension, or vapor pressure
- · Describe the type and properties of one of the four solids
- Draw or use phase diagrams or heating curves to describe phase changes
- Define each of the following terms clearly and completely.
  - (a) condensed state of matter -
  - (b) phase -
  - (c) phase diagram -
  - (d) heating curve -
  - (e) intramolecular force -
  - (f) intermolecular force -
  - (g) metallic bond -
  - (h) hydrogen bond -
  - (i) dipole-dipole force –
  - (j) dispersion force –
  - (k) permanent dipole –
  - (I) temporary dipole -
  - (m) induced dipole -
  - (n) viscosity -
  - (o) surface tension -
  - (p) water -
  - (r) crystalline solid -

- (s) amorphous solid
- (t) adhesion -
- (u) cohesion -
- (v) covalent-network solid -
- (w) phase change -
- (x) vapor pressure -
- (y) heat of fusion -
- (z) heat of vaporization -
- (aa) sublimation -
- (bb) deposition -
- (cc) vaporization -
- (dd) condensation -
- (ee) phase change -
- (ff) freezing -
- (gg) melting -
- (hh) Kinetic-Molecular Theory of Matter –

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**TD Snyder** 

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#### **Table of Contents Review Sheet Chemistry A Study Of Matter**

- 1. Understanding the eBook Review Sheet Chemistry A Study Of Matter
  - The Rise of Digital Reading Review Sheet Chemistry A Study Of Matter
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Review Sheet Chemistry A Study Of Matter
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Review Sheet Chemistry A Study Of Matter
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Review Sheet Chemistry A Study Of Matter
  - Personalized Recommendations
  - Review Sheet Chemistry A Study Of Matter User Reviews and Ratings
  - Review Sheet Chemistry A Study Of Matter and Bestseller Lists
- 5. Accessing Review Sheet Chemistry A Study Of Matter Free and Paid eBooks
  - Review Sheet Chemistry A Study Of Matter Public Domain eBooks
  - Review Sheet Chemistry A Study Of Matter eBook Subscription Services
  - Review Sheet Chemistry A Study Of Matter Budget-Friendly Options
- 6. Navigating Review Sheet Chemistry A Study Of Matter eBook Formats

- o ePub, PDF, MOBI, and More
- Review Sheet Chemistry A Study Of Matter Compatibility with Devices
- Review Sheet Chemistry A Study Of Matter Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Review Sheet Chemistry A Study Of Matter
  - Highlighting and Note-Taking Review Sheet Chemistry A Study Of Matter
  - Interactive Elements Review Sheet Chemistry A Study Of Matter
- 8. Staying Engaged with Review Sheet Chemistry A Study Of Matter
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - o Following Authors and Publishers Review Sheet Chemistry A Study Of Matter
- 9. Balancing eBooks and Physical Books Review Sheet Chemistry A Study Of Matter
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Review Sheet Chemistry A Study Of Matter
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Review Sheet Chemistry A Study Of Matter
  - $\circ\,$  Setting Reading Goals Review Sheet Chemistry A Study Of Matter
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review Sheet Chemistry A Study Of Matter
  - $\circ\,$  Fact-Checking eBook Content of Review Sheet Chemistry A Study Of Matter
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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