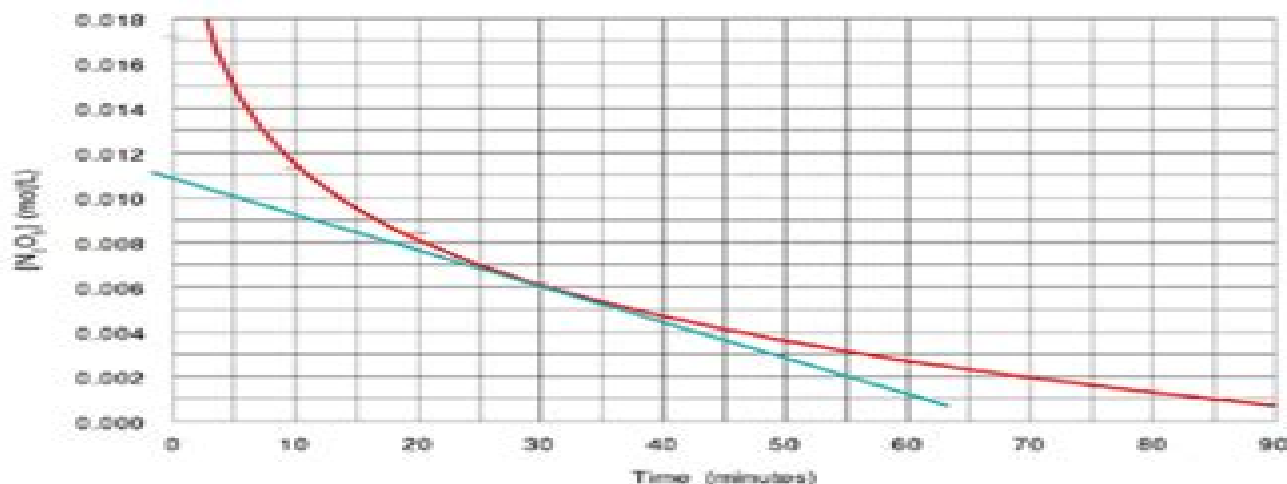


## Kinetics Review Answers

- Pressure, temperature, pH, conductivity, color, mass
- To find the rate we need to calculate the slope of the tangent line at  $t=30$  minutes.



$$\text{rate} = -\frac{\Delta[\text{N}_2\text{O}_5]}{\Delta t} = -\frac{(0.002 - 0.010)}{(55 - 5)} = 0.00016 \text{ mol/L min}$$

- use the rate equation

$$\text{a. } \frac{1}{2} \frac{\Delta[\text{N}_2]}{\Delta t} = \frac{1}{6} \frac{\Delta[\text{H}_2\text{O}]}{\Delta t}$$

$$\frac{1}{2} (0.27) = \frac{1}{6} \frac{\Delta[\text{H}_2\text{O}]}{\Delta t}$$

$$\frac{\Delta[\text{H}_2\text{O}]}{\Delta t} = 0.81 \text{ mol/Ls}$$

$$\text{b. } \frac{1}{2} \frac{\Delta[\text{N}_2]}{\Delta t} = \frac{1}{4} \frac{\Delta[\text{NH}_3]}{\Delta t}$$

$$\frac{1}{2} (0.27) = \frac{1}{4} \frac{\Delta[\text{NH}_3]}{\Delta t}$$

$$\frac{\Delta[\text{NH}_3]}{\Delta t} = 0.54 \text{ mol/Ls}$$

$$\text{c. } \frac{1}{2} \frac{\Delta[\text{N}_2]}{\Delta t} = \frac{1}{3} \frac{\Delta[\text{O}_2]}{\Delta t}$$

$$\frac{1}{2} (0.27) = \frac{1}{3} \frac{\Delta[\text{O}_2]}{\Delta t}$$

$$\frac{\Delta[\text{O}_2]}{\Delta t} = 0.405 \text{ mol/Ls}$$

- Surface area – an increased surface area provides more particles for collisions; more collisions means an increase in reaction rate

Concentration – increased concentration means more particles are present; more particles means more collisions which means an increase in reaction rate

Temperature – an increase in temperature increases the energy of the particles; increased energy means that more of the collisions will have sufficient energy to form the activated complex resulting in an increased reaction rate.

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*Publications* ,1970      **Corrosion** Ambrish Singh,2020-10-21 Corrosion is a global threat and a burning topic for new and innovative research Corrosion causes shut downs economic losses delays failures accidents losses of life and losses in productivity Wherever metal is there corrosion will occur this is a general concept as not many protection methods are available to mitigate corrosion The available methods can only delay the process but cannot stop or protect the metal completely So there is always a need for good research and inventions in this field This book includes the recent research work done in the field of corrosion The chapters are written by reputed authors in the field of corrosion and have been reviewed extensively before acceptance The chapters focus on different aspects of corrosion to provide readers with a good idea of the overall process The diversification of the chapters will keep the readers interested and motivated for new innovations in the field of corrosion It will be very useful to scholars academicians researchers and industrialists

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