

Review #5: Chemical Equations and Stoichiometry

- Define limiting factor. Why is it significant?
The limiting factor is the reactant that runs out first in a chemical reaction. The limiting factor is significant because once it is gone, the reaction stops. It is the amount of the limiting factor that determines how much product can be produced.
- Find the simplest formula for the compound with composition:
 - 38.7% carbon, 9.7% hydrogen and 51.6% oxygen (CH_3O)
 - 82.4% nitrogen and 17.6% hydrogen (NH_3)
- A certain compound is 40.0% carbon, 6.7% hydrogen and 53.3% oxygen by weight. One mole of this substance weighs 180 grams. What is the molecular formula of the compound? ($\text{C}_6\text{H}_{12}\text{O}_6$)
- Balance these expressions. Type of Reaction
 - $4 \text{ Al} + 3 \text{ O}_2 \rightarrow 2 \text{ Al}_2\text{O}_3$ synthesis
 - $2 \text{ C}_2\text{H}_6 + 7 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 6 \text{ H}_2\text{O}$ combustion
 - $3 \text{ MgO} + 2 \text{ H}_3\text{PO}_4 \rightarrow \text{Mg}_3(\text{PO}_4)_2 + 3 \text{ H}_2\text{O}$ double displacement
 - $2 \text{ PbO}_2 \rightarrow 2 \text{ PbO} + \text{O}_2$ decomposition
 - $\text{SiO}_2 + 4 \text{ HF} \rightarrow \text{SiF}_4 + 2 \text{ H}_2\text{O}$ double displacement
 - $2 \text{ C}_{10}\text{H}_{22} + 31 \text{ O}_2 \rightarrow 20 \text{ CO}_2 + 22 \text{ H}_2\text{O}$ combustion
 - $\text{Mg} + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2 + \text{MgSO}_4$ single displacement
 - $\text{Sb}_2\text{S}_3 + 12 \text{ HCl} \rightarrow 2 \text{ H}_3\text{SbCl}_6 + 3 \text{ H}_2\text{S}$ Can't classify
- For the reaction: $2 \text{ Fe}_2\text{O}_3 + 3 \text{ C} \rightarrow 4 \text{ Fe} + 3 \text{ CO}_2$
What mass of iron will be produced when 50.0 g of iron(III) oxide react with an excess of carbon? (35.0 g)
- For the reaction: $2 \text{ KClO}_{3(s)} \rightarrow 2 \text{ KCl}_{(s)} + 3 \text{ O}_{2(g)}$
 - What volume of oxygen gas, measured at STP, is formed by the reaction of 75.0 g of potassium chlorate? (20.6 L)
 - How many moles of potassium chlorate must react in order to form 2.5 g of potassium chloride? (0.034 mol)
- For the reaction: $6 \text{ CO}_{2(g)} + 6 \text{ H}_2\text{O}_{(g)} \rightarrow \text{C}_6\text{H}_{12}\text{O}_{6(s)} + 6 \text{ O}_{2(g)}$
 - Which substance is in excess when 100 g of carbon dioxide and 50.0 g of water are reacted? (water)
 - How much glucose will be produced by the reaction in part (a)? (68.2 g)
- Octane burns according to the reaction: $2 \text{ C}_8\text{H}_{18(l)} + 25 \text{ O}_{2(g)} \rightarrow 16 \text{ CO}_{2(g)} + 18 \text{ H}_2\text{O}_{(g)}$
 - What mass of carbon dioxide will be produced when 180 g of octane are completely burned? (555 g)
 - What volume of carbon dioxide, at STP, will be produced when 62.7 g of oxygen are completely reacted? (28.1 L)

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Table of Contents Review Answers Chemical Reactiona

1. Understanding the eBook Review Answers Chemical Reactiona
 - The Rise of Digital Reading Review Answers Chemical Reactiona
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Answers Chemical Reactiona
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Answers Chemical Reactiona
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review Answers Chemical Reactiona
 - Personalized Recommendations
 - Review Answers Chemical Reactiona User Reviews and Ratings
 - Review Answers Chemical Reactiona and Bestseller Lists
5. Accessing Review Answers Chemical Reactiona Free and Paid eBooks
 - Review Answers Chemical Reactiona Public Domain eBooks
 - Review Answers Chemical Reactiona eBook Subscription Services
 - Review Answers Chemical Reactiona Budget-Friendly Options
6. Navigating Review Answers Chemical Reactiona eBook Formats
 - ePub, PDF, MOBI, and More
 - Review Answers Chemical Reactiona Compatibility with Devices
 - Review Answers Chemical Reactiona Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review Answers Chemical Reactiona
 - Highlighting and Note-Taking Review Answers Chemical Reactiona
 - Interactive Elements Review Answers Chemical Reactiona

8. Staying Engaged with Review Answers Chemical Reactiona
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Answers Chemical Reactiona
9. Balancing eBooks and Physical Books Review Answers Chemical Reactiona
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Answers Chemical Reactiona
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Review Answers Chemical Reactiona
 - Setting Reading Goals Review Answers Chemical Reactiona
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Review Answers Chemical Reactiona
 - Fact-Checking eBook Content of Review Answers Chemical Reactiona
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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