



Use the equation above to answer the following three questions.
Starting with 10.0 mol of Al, determine the theoretical yield in grams of H_2 .

$$\frac{10 \text{ mol Al}}{2 \text{ mol Al}} \times \frac{3 \text{ mol H}_2}{1 \text{ mol H}_2} = 30 \text{ g H}_2$$

2. Use the equation above to answer the following question.
If the actual yield is 28.2 g of H_2 , calculate the percent yield?

$$\% \text{ yield} = \frac{\text{Actual}}{\text{Theoretical}} \times 100$$

$$\frac{28.2 \text{ g H}_2}{30 \text{ g H}_2} \times 100 = \boxed{94\%}$$

3. Use equation above to answer the following question.
How many moles of H_2SO_4 are needed to completely react with 15.0 mol of Al?

$$\frac{15 \text{ mol Al}}{2 \text{ mol Al}} \times \frac{3 \text{ mol H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} = \boxed{22.5 \text{ mol H}_2\text{SO}_4}$$

4. For every 9 moles of NiO there are how many moles of Al. (must balance)



$$\frac{9 \text{ mol NiO}}{3 \text{ mol NiO}} \times \frac{2 \text{ mol Al}}{1 \text{ mol Al}} = \boxed{6 \text{ mol Al}}$$

5. How many liters of oxygen are required to react completely with 3.6 liters of hydrogen to produce water? (hint you need to write the balanced equation for the formation of water to get the mole ratio)



$$\frac{3.6 \text{ L H}_2}{22.4 \text{ L H}_2} \times \frac{1 \text{ mol H}_2}{2 \text{ mol H}_2} \times \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{22.4 \text{ L O}_2}{1 \text{ mol O}_2} = \boxed{1.8 \text{ L O}_2}$$

6. If 17.05 grams of ethene (C_2H_4) burns in the presence of 14.0 grams Oxygen, what is the Limiting reactant and the excess? (Go to grams of H_2O)



$$\frac{17.05 \text{ g C}_2\text{H}_4}{28 \text{ g C}_2\text{H}_4} \times \frac{1 \text{ mol C}_2\text{H}_4}{1 \text{ mol C}_2\text{H}_4} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol C}_2\text{H}_4} \times \frac{18 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 21.9 \text{ g H}_2\text{O}$$

$$\frac{14 \text{ g O}_2}{32 \text{ g O}_2} \times \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{2 \text{ mol H}_2\text{O}}{3 \text{ mol O}_2} \times \frac{18 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{5.25 \text{ g H}_2\text{O}}$$

Smaller Amount

Limiting: 14 g O₂ Excess: 17.05 g C₂H₄

Review Chemistry Honors Unit Stoichiometry

**Martin Van Buren McGill, G. M.
Bradbury**



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