

Chemistry 1020 Final Exam Review

1. A 25.0 mL sample of 1.50 M sodium chlorate (NaClO) is titrated with 1.00 M HCl . The K_a of HClO is 2.9×10^{-8} .

Weak Acid

(a) What is the pH after 20.0 mL of HCl has been added?



$n_{\text{ClO}^-} = (1.50\text{M})(0.0250\text{L}) = 0.0375\text{mol}$ ClO^- is in excess

$n_{\text{H}_3\text{O}^+} = (1.00\text{M})(0.0200\text{L}) = 0.0200\text{mol}$ all H_3O^+ reacts

buffer: still we are between start and equivalence point both ClO^- and HClO are in solution

$\text{pH} = \text{p}K_a + \log \frac{[\text{ClO}^-]}{[\text{HClO}]}$

$K_a = \text{HClO}$
 $K_a = \text{ClO}^-$

$n_{\text{ClO}^-} \text{ remaining} = n_{\text{ClO}^-} - n_{\text{H}_3\text{O}^+} = 0.0375\text{mol} - 0.0200\text{mol} = 0.0175\text{mol}$
 $n_{\text{HClO}} = n_{\text{H}_3\text{O}^+} = 0.0200\text{mol}$ $\text{pH} = -\log(2.9 \times 10^{-8}) + \log\left(\frac{0.0175}{0.0200}\right) = 7.48$

(b) What is the pH at the equivalence point?

eq. point: all ClO^- consumed. n_{HClO} added = n_{ClO^-} from 3rd



$n_{\text{HClO}} = n_{\text{ClO}^-} \text{ added}$

$\frac{0.0375\text{mol}}{1.00\text{M}} = 0.0375\text{L}$

$[\text{HClO}] = \frac{n_{\text{HClO}}}{V_{\text{total}}} = \frac{0.0375\text{mol}}{0.025\text{L} + 0.0375\text{L}} = 0.600\text{M}$

$\text{pH} = -\log(0.600) = -\log(6 \times 10^{-1}) = 0.22$

(c) What is the pH after 40.0 mL of HCl have been added?

eq. V of HCl was 37.5 mL

V of HCl after eq. point = $40.0\text{mL} - 37.5\text{mL} = 2.5\text{mL}$
 $= 0.0025\text{L}$

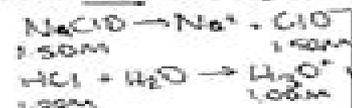
$n_{\text{H}_3\text{O}^+} \text{ in solution} = (0.0025\text{L})(1.00\text{M}) = 0.0025\text{mol} = n_{\text{H}_3\text{O}^+}$

$[\text{H}_3\text{O}^+] = \frac{n_{\text{H}_3\text{O}^+}}{V_{\text{total}}} = \frac{0.0025\text{mol}}{0.025\text{L}} = 0.100\text{M}$

$\text{pH} = -\log(0.100) = -\log(10^{-1}) = 1.00$

Weak Acid

Strong Acid



C-17

Review Problems For Regular Chemistry Final

Peipei Pang



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