

AP® CHEMISTRY 2005 EXAMINATIONS CHEMISTRY SECTION

Duration 45
min

As sample of nitrogen is subjected to constant pressure, a plot of its volume (in liters) as a function of the pressure is approximately given by the graph below.



The following PV data are consistent with the pressure-volume relationship shown in the graph above. These data represent the behavior of the gas at the temperature shown.

Temperature (K)

- 1.00 mol of the gas at 1.00 atm and 1.00 L
- 1.00 mol of the gas at 2.00 atm and 2.00 L
- 1.00 mol of the gas at 3.00 atm and 3.00 L
- 1.00 mol of the gas at 4.00 atm and 4.00 L
- 1.00 mol of the gas at 5.00 atm and 5.00 L

1 point is earned for each correct response.

1. The gas is heated at constant pressure. The volume of the gas increases by 1.00 L.

$$\Delta V = 1.00 \text{ L} \quad \Delta P = 0 \text{ atm} \quad \Delta T = ? \text{ K} \quad \Delta n = 0$$

2. The gas is cooled at constant pressure. The volume of the gas decreases by 1.00 L.

- 1.00 mol of the gas at 1.00 atm and 1.00 L
- 1.00 mol of the gas at 2.00 atm and 2.00 L
- 1.00 mol of the gas at 3.00 atm and 3.00 L
- 1.00 mol of the gas at 4.00 atm and 4.00 L
- 1.00 mol of the gas at 5.00 atm and 5.00 L

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Leon P. Bignold



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