

14-1 Review and Reinforcement (continued)

Complete each of the sentences below.

15. The stronger the intermolecular forces in a liquid, the _____ its boiling point.
16. The state of a substance at room temperature depends on _____.
17. Water molecules in an ice cube are held together by _____ forces or, more specifically, _____ bonds.

On the line at the left, write the letter of the term that matches each description below. Each choice can be used once, more than once, or not at all.

a. dispersion

b. dipole-dipole

c. hydrogen bonding

- _____ 18. the force in between SO_2 molecules
- _____ 19. the force that accounts for HF being a liquid while H_2 and F_2 are gases
- _____ 20. the force that depends on temporary, induced dipoles
- _____ 21. the only intermolecular force that exists in noble gases
- _____ 22. the attractive force between neighboring dipoles
- _____ 23. the force that arises because of the large difference in electronegativities in the N-H bonds

Answer the following question as directed.

24. Which of the following statements could be used to explain why water has an unusually high boiling point? Make a checkmark on the line for all that apply.
- _____ a. Water molecules have a strong intermolecular force.
- _____ b. There are no nitrogen atoms in water.
- _____ c. Oxygen is more electronegative than hydrogen.
- _____ d. Water is an angular molecule.
- _____ e. Oxygen atoms are smaller than hydrogen atoms.
- _____ f. Water does not contain metallic bonds.
- _____ g. There is a large difference in electronegativity between oxygen and hydrogen.

Review And Reinforcement Answers Astronomy

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Table of Contents Review And Reinforcement Answers Astronomy

1. Understanding the eBook Review And Reinforcement Answers Astronomy
 - The Rise of Digital Reading Review And Reinforcement Answers Astronomy
 - Advantages of eBooks Over Traditional Books
2. Identifying Review And Reinforcement Answers Astronomy
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review And Reinforcement Answers Astronomy
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review And Reinforcement Answers Astronomy
 - Personalized Recommendations
 - Review And Reinforcement Answers Astronomy User Reviews and Ratings
 - Review And Reinforcement Answers Astronomy and Bestseller Lists

5. Accessing Review And Reinforcement Answers Astronomy Free and Paid eBooks
 - Review And Reinforcement Answers Astronomy Public Domain eBooks
 - Review And Reinforcement Answers Astronomy eBook Subscription Services
 - Review And Reinforcement Answers Astronomy Budget-Friendly Options
6. Navigating Review And Reinforcement Answers Astronomy eBook Formats
 - ePub, PDF, MOBI, and More
 - Review And Reinforcement Answers Astronomy Compatibility with Devices
 - Review And Reinforcement Answers Astronomy Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review And Reinforcement Answers Astronomy
 - Highlighting and Note-Taking Review And Reinforcement Answers Astronomy
 - Interactive Elements Review And Reinforcement Answers Astronomy
8. Staying Engaged with Review And Reinforcement Answers Astronomy
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review And Reinforcement Answers Astronomy
9. Balancing eBooks and Physical Books Review And Reinforcement Answers Astronomy
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review And Reinforcement Answers Astronomy
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Review And Reinforcement Answers Astronomy
 - Setting Reading Goals Review And Reinforcement Answers Astronomy
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Review And Reinforcement Answers Astronomy
 - Fact-Checking eBook Content of Review And Reinforcement Answers Astronomy
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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