


Ideal Gases

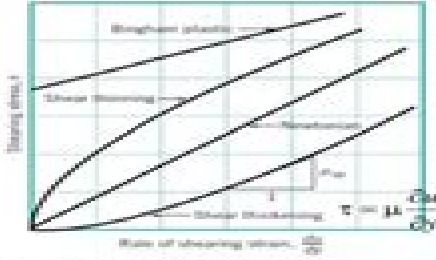
- From chemistry: $PV = nRT$ (V is volume)
 - $n = m / M$ is the number of moles
 - for mass in kg n is in kilogram moles (kmol); for mass in lb_m, n is in pound moles (lbmol)
 - $R = 8.31447$ kJ/kmol-K = 10.7316 psia-ft³ / lbmol-R is universal gas constant
 - $R = R/M$ is engineering gas constant that is different for each gas
 - Real gases like ideal gases at low pressures
 - $P = nRT / V = (m/M)RT / V = (m/V)(R/M)T$
 $P = \rho RT$




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Viscosity


Newtonian and non-Newtonian
Variation of shear stress with rate of shearing strain for several types of fluids, including common non-Newtonian fluids.






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Surface Tension Rise



- Vertical force balance: $\gamma \pi R^2 h = 2 \pi R \sigma \cos \theta$
 - Surface tension depends on fluid and temperature, wetting angle, θ , depends on fluid and surface


$$h = \frac{2 \sigma \cos \theta}{\gamma R}$$



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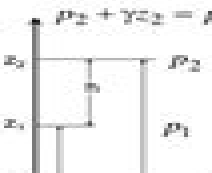
Pressure Relations

- Pressure is a scalar
- The force exerted by a pressure is the same in all directions
- Want to see how pressure changes in a static (nonmoving fluid)
- Look at balance of pressure force and fluid weight over a differential volume element, $\Delta x \Delta y \Delta z$




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Incompressible Fluid



- Which pressure is higher?
 $p_1 = p_2 + \gamma(z_2 - z_1)$
 $p_1 = p_2 + \gamma h > p_2$
- Pressure increases with depth



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Gage and Absolute Pressure





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M Woodhall



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