

10.1

THE NATURE OF GASES

SECTION REVIEW

Objectives

- Describe the motion of gas particles according to the kinetic theory
- Interpret gas pressures in terms of kinetic theory

Key Terms

- kinetic energy
- kinetic theory
- gas pressure
- vacuum
- atmospheric pressure
- barometric pressure
- perfect (P₀)
- standard atmospheric (atm)

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are highlighted in this section. Write clearly over the completed work a term, short phrase, or sentence.

- The kinetic theory describes the _____ of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly _____, that the particles of gas are _____, move _____ of each other, and are in constant _____ motion. Furthermore, the _____ between particles are elastic so that the total _____ remains constant. The pressure of a gas results from the countless collisions of the gas particles with an object. Barometers are used to measure _____ pressure. Standard conditions are defined as a temperature of _____ and a pressure of _____.
- _____
 - _____
 - _____
 - _____
 - _____
 - _____
 - _____
 - _____
 - _____
 - _____

Part B True-False

Classify each of these statements as true (always true), AT (sometimes true), ST (sometimes true), or F (false).

- Atmospheric pressure is 760 mm Hg.
- The SI unit of pressure is the pascal.
- Atmospheric pressure increases as you climb a mountain because the air layer that compresses the air gets thicker with height.

Review Gases Section 4 Answers

Monica M. Tighe, Marg Brown



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