Answers to the Gases Review Sheet

 A sample of air has a volume of 2.25 cm³ when its temperature is 298 K. If the temperature is increased to 373 K without changing the pressure, what is the new volume of the sample in air ?

$$V_1 = 2.25 \text{ cm}^3$$
 $V_1 = V_2$ $V_2 = V_1 T_2$ $V_3 = V_2 T_1$ $V_4 = V_2 T_2$ $V_5 = V_1 T_2$ $V_7 = V_1 T_2$ $V_8 = V_1 T_2$ $V_9 = V_1$

6. As the temperature of a sample of nitrogen increases from 273 K, the volume of the sample changes from 275 mL to 325 mL. Assuming the pressure is constant, what is the final temperature of the gas ?

$$V_1 = 275 \text{ mL}$$
 $V_1 = V_2$ $T_2 = \frac{T_1V_2}{V_1}$
 $T_1 = 273 \text{ K}$ $T_1 = V_2$ V_1
 $T_2 = x$
 $V_2 = 325 \text{ mL}$ Chark's
 $T_2 = \frac{325 \text{ mL}(273 \text{ K})}{275 \text{ mL}} = 323 \text{ K}$

The pressure of a 5.71 L sample of neon gas is 23.4 kPa. Calculate the new pressure when the volume becomes 3.40 L.

$$V_1 = 5.71 \text{ L}$$
 $P_1 = 23.4 \text{ kPa}$
 $P_2 = X$
 $V_2 = 3.40 \text{ L}$
 $P_2 = 5.71 \text{ L}(23.4 \text{ kPa}) = 39.3 \text{ kPa} = 0.388 \text{ atm}$
 3.40 L

Review Section Gases Answers

Tobias Bleicker

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