



## Unit: Chemical Bonding Answer Key “Covalent Bonding” - WS #3

**Directions:** Please answer each question using your notes from today’s class. ☺

1. What is the difference between a molecular formula, structural formula and an electron dot formula? Give an example of each.

*A molecular formula gives the chemical make-up of one molecule. It does not show atom-to-atom connectivity. Example:  $C_3H_8$*

*A structural formula uses lines instead of dots to show shared pairs of electrons.*

*Example:  $H-H$*

*An electron dot formula uses dots instead of lines to show shared pairs of electrons.*

*Example:  $H:H$*

2. What is the difference between single, double, and triple covalent bonds? How many shared electrons are in each?

*The difference between the types of bonds is: a single bond is two shared electrons, a double is four shared electrons, and a triple is six shared electrons.*

3. Draw the electron dot and structural formulas for the following molecules.

Substance	Rough copy (work!)	Electron dot	Structural
a) $Cl_2$		<i>Note: The electron dot structures have dots in between each atom instead of lines.</i>	$\text{:}\ddot{\text{Cl}}-\ddot{\text{Cl}}\text{:}$
b) $H_2S$			$\begin{array}{c} \text{:}\ddot{\text{S}}\text{:} \\   \\ \text{H}-\end{array}$
c) $BrI$			$\text{:}\ddot{\text{I}}-\ddot{\text{Br}}\text{:}$

# **Review Answers For Covalent Bonding**

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Viselli**



## **Review Answers For Covalent Bonding:**

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Ebbing,2005-03-17 Provides solutions to odd numbered Practice Problems General Problems and Cumulative Skills Problems plus answers to Review Questions      *Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J.

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