

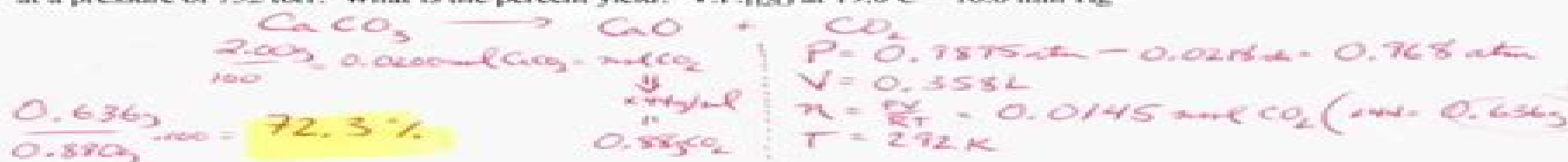
7. A student's automobile uses 1/3 gallon (1160 grams) of octane on each round trip to and from school. What volume of  $\text{CO}_2$  gas is produced at  $-15^\circ\text{C}$  and 29.55 in Hg?



8. Given:  $2\text{NH}_4\text{Cl}_{(\text{aq})} + \text{Ca}(\text{OH})_{2(\text{aq})} \rightarrow \text{CaCl}_{2(\text{aq})} + 2\text{NH}_{3(\text{g})} + 2\text{H}_2\text{O}_{(\text{l})}$ . What volume of ammonia gas is produced at  $72.0^\circ\text{C}$  and 3.55 atm when 6.20g of ammonium chloride reacts with 9.00g of calcium hydroxide?



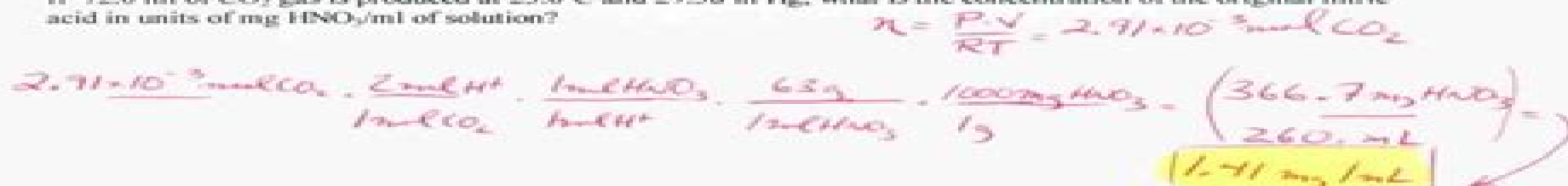
9. From the decomposition of 2.00grams of  $\text{CaCO}_3$  a student collects 358 ml of  $\text{CO}_2$  gas over water at  $19.0^\circ\text{C}$  and at a pressure of 752 torr. What is the percent yield? V.P.<sub>19.0</sub> at  $19.0^\circ\text{C}$  = 16.6 mm Hg



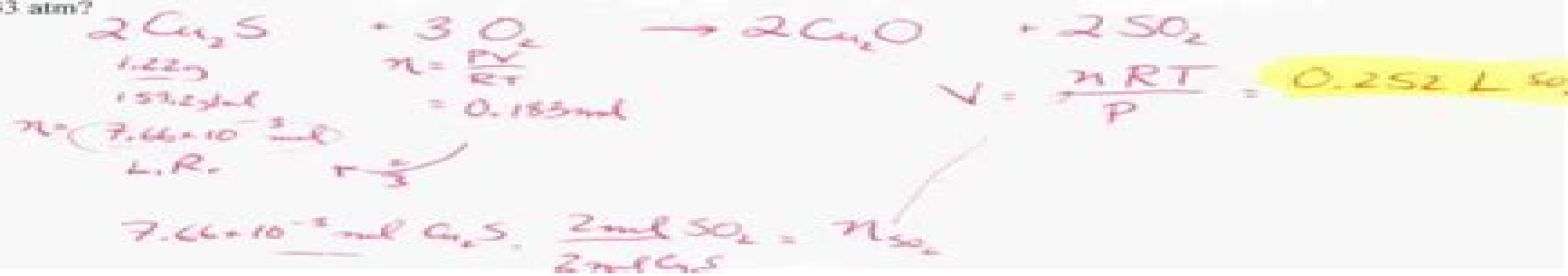
10. 260.ml of a solution of nitric acid,  $\text{HNO}_3$ , is analyzed with an excess of  $\text{CaCO}_3$  by the reaction:



If 72.0 ml of  $\text{CO}_2$  gas is produced at  $23.0^\circ\text{C}$  and 29.38 in Hg, what is the concentration of the original nitric acid in units of mg  $\text{HNO}_3$ /ml of solution?



11. Consider the following unbalanced equation:  $\text{Cu}_2\text{S}_{(\text{s})} + \text{O}_{2(\text{g})} \rightarrow \text{Cu}_2\text{O}_{(\text{s})} + \text{SO}_{2(\text{g})}$ . If 1.22 grams of  $\text{Cu}_2\text{S}$  reacted with 4.52 liters of  $\text{O}_2$  gas at  $28^\circ\text{C}$  and 1.0 atm what volume of  $\text{SO}_2$  gas would be formed at  $100.^\circ\text{C}$  and 0.933 atm?



# Sample Problems H Gas Stoichiometry Answers

**Stephen DeMeo**



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*Pp/Chemistry* BarCharts, Inc., 2008-06-18

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