

balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds.

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20 \\ + & & \\ \hline & & 63.62 \rightarrow 64 \end{array}$$

Test Modern Chemistry Solutions Answers

Everett Brackin Tucker, H. L. Lambert



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Laboratory Experiments to Accompany "Modern Chemistry," Fredus Nelson Peters, 1902 Modern Chemistry Nicholas D. Tzimopoulos, 1990 Modern Inorganic Chemistry Joseph William Mellor, 1917 **Modern Cereal Chemistry** Douglas William Kent-Jones, 1924 Lab Experiments for Modern Chemistry Tzimopoulo, 1990 The Handy Chemistry Answer Book Justin P. Lomont, Ian C. Stewart, 2013-10-01

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Experimental Chemistry George W. Jr. Latimer,2012-12-02 Modern Experimental Chemistry provides techniques of

qualitative analysis that reinforce experiments on ionic equilibria. This book includes the determination of water in hydrated salts, identification of an organic compound after determining its molecular weight, and nonaqueous titration of a salt of a weak acid. The calculation of chemical stoichiometry, calculation of thermodynamic properties by determining the change in equilibrium with temperature, and chromium chemistry are also covered. This compilation contains enough experiments for classes which have six hours of laboratory two 3 hour meetings per week to last two semesters. This publication is intended for chemistry students as an introductory manual to chemistry laboratory.

Chemistry Eileen Ramsden, 2001. This third edition of Key Science Chemistry has been fully revised to meet the requirements of all 2001 GCSE specifications. It is aimed at middle ability students but contains enough material for high achievers. Topics are clearly differentiated between core material for GCSE science Double Award Single Award and extension material for GCSE science chemistry.

Fundamentals of Chemistry: A Modern Introduction Frank Brescia, 2012-12-02. Fundamentals of Chemistry Third Edition introduces the reader to the fundamentals of chemistry including the properties of gases, atomic and molecular weights, and the first and second laws of thermodynamics. Chemical equations and chemical arithmetic are also discussed along with the structure of atoms, chemical periodicity, types of chemical bonds, and condensed states of matter. This book is comprised of 26 chapters and begins with a historical overview of chemistry and some terms which are part of the language of chemists. Separation and purification are covered in the first chapter while the following chapters focus on atomic and molecular weights, stoichiometry, the structure of atoms, and types of chemical bonds. The molecular orbital (MO) theory of bonding, galvanic cells, and chemical thermodynamics are considered next. Separate chapters are devoted to MO theory of covalent and metallic bonding, orbital hybridization, intermolecular forces, acids and bases, ionic equilibrium, calculations, and polymers and biochemicals. This monograph is intended for chemistry students.

School Science and Mathematics, 1926. *The Journal of Arkansas Education*. Everett Brackin, Tucker, H. L. Lambert, 1929.

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