

# CHAPTER 19

## ELECTROCHEMISTRY

### Problem Categories

*Biological:* 19.66, 19.68, 19.126.

*Conceptual:* 19.67, 19.79, 19.85, 19.94, 19.103, 19.106, 19.108, 19.128.

*Descriptive:* 19.13, 19.14, 19.17, 19.18, 19.46a, 19.52a, 19.61, 19.77, 19.87, 19.93, 19.99, 19.100, 19.101, 19.111, 19.113, 19.123.

*Environmental:* 19.63.

*Industrial:* 19.48, 19.56, 19.89, 19.109, 19.112, 19.120.

*Organic:* 19.38.

### Difficulty Level

*Easy:* 19.11, 19.12, 19.13, 19.14, 19.16, 19.17, 19.18, 19.22, 19.45, 19.63, 19.67, 19.79, 19.91, 19.97.

*Medium:* 19.1, 19.2, 19.15, 19.21, 19.23, 19.24, 19.25, 19.26, 19.29, 19.30, 19.31, 19.32, 19.33, 19.34, 19.38, 19.46, 19.47, 19.48, 19.49, 19.50, 19.51, 19.52, 19.53, 19.55, 19.57, 19.58, 19.59, 19.60, 19.61, 19.62, 19.64, 19.65, 19.66, 19.69, 19.70, 19.72, 19.73, 19.74, 19.75, 19.77, 19.81, 19.82, 19.83, 19.84, 19.85, 19.86, 19.87, 19.89, 19.93, 19.94, 19.96, 19.98, 19.99, 19.100, 19.103, 19.104, 19.106, 19.107, 19.108, 19.109, 19.110, 19.113, 19.114, 19.115, 19.116, 19.117, 19.122, 19.124, 19.126.

*Difficult:* 19.37, 19.54, 19.56, 19.68, 19.71, 19.76, 19.78, 19.80, 19.88, 19.90, 19.92, 19.95, 19.101, 19.102, 19.105, 19.111, 19.112, 19.118, 19.119, 19.120, 19.121, 19.123, 19.125, 19.127, 19.128.

**19.1** We follow the steps as described in detail in Section 19.1 of the text.

(a) The problem is given in ionic form, so combining Steps 1 and 2, the half-reactions are:



**Step 3:** We balance each half-reaction for number and type of atoms and charges.

The *oxidation half-reaction* is already balanced for Fe atoms. There are three net positive charges on the right and two net positive charges on the left, we add one electron to the right side to balance the charge.



*Reduction half-reaction:* we add one  $\text{H}_2\text{O}$  to the right-hand side of the equation to balance the O atoms.



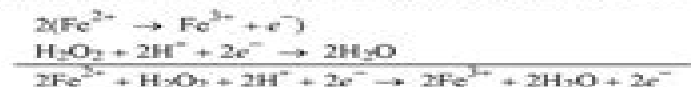
To balance the H atoms, we add  $2\text{H}^{+}$  to the left-hand side.



There are two net positive charges on the left, so we add two electrons to the same side to balance the charge.



**Step 4:** We now add the oxidation and reduction half-reactions to give the overall reaction. In order to equalize the number of electrons, we need to multiply the oxidation half-reaction by 2.



# Test Exam Of Electrochemistry Chap 19

**RM Cervero**



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