Precipitate Reactions: Chemical "Drop-Outs"

Many chemical reactions take place in solution. In this experiment, you will see some dissolved chemicals react when solutions are combined and drop out of solution or precipitate. In a chemical reaction, substances that react to form new substances are called reactants; new substances that are produced are called products. Substances that "fall out" of solution as solids, denoted not called precipitates (ppt).

Goals:

- To learn about double replacement reactions.
- To learn solubility rules.
- To learn small-scale lab techniques.
- To write balanced chemical equations to describe reactions performed in lab.

Materials Needed:

- 6 small scale pipettes (empty, barium chloride, magnesium nitrate, copper (II) sulfate, potassium hydroxide, and sodium carbonate)
- reaction surface (Make sure the surface is clean, but DO NOT USE SOAP.)

Procedure:

- 1. Obtain materials.
- Put 3 drops of sodium carbonate on the reaction surface in A. Add 3 drops of barium chloride solution. Be careful to not cross contaminate the tip of the pipets.
- Record your observations in the data chart. Pay special attention to the color of each
 precipitate that forms, as well as its consistency (grainy, clumpy, etc.). Use a Q-tip if
 needed to assist with your observations. Write "No Rxn" when no precipitate forms.
- Make all of the remaining combinations listed in the data chart on the reaction surface in boxes B-J. Use 3 drops of each solution as you did in step 2. Record your observations in the data chart as soon as you make a combination.
- Return all pipets bulb down to the proper container and clean the reaction surface with a paper towel.

Post-labe

- Complete the word equations for the reaction under part 2 in the equation chart.
- 7. Write the chemical equation for part 1 in the equation chart for each combination.

For your equations from part 1 in the equation chart:

- If a precipitate formed, use the solubility chart to determine which of the products formed is the precipitate (use the solubility table). Circle the precipitate in the equation and label with _{int}. Label the other product with _{int}.
- If no precipitate is formed, label each product with the number in the pair column.
- Answer post-lab questions.

Precipitate Lab Answers

Olga I. Agapova

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