

# 13.1 THE NATURE OF GASES

P.1

18 pts  
 - 2 pt if late  
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## Section Review

### Objectives

- Describe the assumptions of the kinetic theory as it applies to gases
- Interpret gas pressure in terms of kinetic theory
- Define the relationship between Kelvin temperature and average kinetic energy

### Vocabulary

- kinetic energy
- kinetic theory
- gas pressure
- vacuum
- atmospheric pressure
- barometer
- pascal (Pa)
- standard atmosphere (atm)

### Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The kinetic theory describes the 1 of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly 2, that the particles of gas are relatively 3, move 4 of each other, and are in constant 5 motion. The 6 between particles are perfectly elastic so that the total 7 remains constant. Gas pressure results from the simultaneous collisions of billions of particles with an object. Barometers are used to measure 8 pressure. Standard conditions are defined as a temperature of 9 and a pressure of 10.

1. motion
2. empty space
3. far apart
4. independently
5. random / rapid
6. collisions
7. Kinetic energy
8. atmospheric
9. 0°C / 273 K
10. 1 atm / 101.3 kPa

### Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- ST 11. Atmospheric pressure is 760 mm Hg.  
AT 12. The SI unit of pressure is the pascal.

↳ a multiple of kilo pascal

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**LM Reder**



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## **Table of Contents Prentice Hall Chemistry 13 1 Worksheet Answers**

1. Understanding the eBook Prentice Hall Chemistry 13 1 Worksheet Answers
  - The Rise of Digital Reading Prentice Hall Chemistry 13 1 Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry 13 1 Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Prentice Hall Chemistry 13 1 Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry 13 1 Worksheet Answers
  - Personalized Recommendations
  - Prentice Hall Chemistry 13 1 Worksheet Answers User Reviews and Ratings
  - Prentice Hall Chemistry 13 1 Worksheet Answers and Bestseller Lists
5. Accessing Prentice Hall Chemistry 13 1 Worksheet Answers Free and Paid eBooks
  - Prentice Hall Chemistry 13 1 Worksheet Answers Public Domain eBooks
  - Prentice Hall Chemistry 13 1 Worksheet Answers eBook Subscription Services
  - Prentice Hall Chemistry 13 1 Worksheet Answers Budget-Friendly Options
6. Navigating Prentice Hall Chemistry 13 1 Worksheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Prentice Hall Chemistry 13 1 Worksheet Answers Compatibility with Devices
  - Prentice Hall Chemistry 13 1 Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry 13 1 Worksheet Answers
  - Highlighting and Note-Taking Prentice Hall Chemistry 13 1 Worksheet Answers
  - Interactive Elements Prentice Hall Chemistry 13 1 Worksheet Answers
8. Staying Engaged with Prentice Hall Chemistry 13 1 Worksheet Answers

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Prentice Hall Chemistry 13 1 Worksheet Answers
9. Balancing eBooks and Physical Books Prentice Hall Chemistry 13 1 Worksheet Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Prentice Hall Chemistry 13 1 Worksheet Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry 13 1 Worksheet Answers
- Setting Reading Goals Prentice Hall Chemistry 13 1 Worksheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry 13 1 Worksheet Answers
- Fact-Checking eBook Content of Prentice Hall Chemistry 13 1 Worksheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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