

FLAME TESTS

INTRODUCTION

According to the Bohr theory of the atom, electrons may occupy only specific energy levels. When an atom absorbs sufficient energy, an electron can "jump" to a higher energy level. Higher energy levels tend to be less stable, however, and if a lower energy level is available, the electron will "fall" back, giving off energy in the process. The difference in energies between the two levels is emitted in the form of a photon of electromagnetic radiation. The energy of each photon is described by the equation $E = h\nu$, where h is Planck's constant and ν is the frequency of the radiation. If the wavelength of the released photon is between 400 and 700 nm, the energy is emitted as visible light. The color of the light depends on the specific energy change that is taking place.

White light is a continuous spectrum in which all wavelengths of visible light are present. An excited atom, however, produces one or more specific lines in its spectrum, corresponding to the specific changes in energy levels of its electrons. Because each element has a distinct electron configuration, each has a unique line spectrum.

Flame tests are a quick method of producing the characteristic colors of metallic ions, the loosely-held electrons of a metal are easily excited in the flame of a lab burner. The emission of energy in the visible portion of the spectrum as those electrons return to lower energy levels produces a colored flame, the color is a combination of the wavelengths of each transition, and may be used to determine the identity of the ion.

In this investigation you will perform flame tests on seven metallic ions, then use your results to determine the identity of several unknowns.

PRE-LAB DISCUSSION

Read the entire laboratory investigation and the relevant pages of your textbook, then answer the questions that follow.

1. Write out the electron configurations for each of the metallic ions to be tested in this investigation.

Ba²⁺ _____
Cu²⁺ _____
Li⁺ _____
K⁺ _____
Sr²⁺ _____
Ca²⁺ _____
Na⁺ _____

2. What does a flame test indicate about the energy changes taking place among the electrons in a metallic ion? _____

Prentice Hall Flame Test Lab Answers

Raffaela Di Napoli



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Prentice Hall Exploring Earth Science ,1997 **Prentice Hall Earth Science** ,1988 Chemistry Eugene LeMay, Jr.,Herbert Beall,Karen M. Robblee,Douglas C. Brower,Prentice-Hall Staff,2002-02 **Matter** Prentice-Hall Staff,1994

Chemistry Dorin,Prentice Hall Dictionary Editors,1992 *Resources for Teaching Middle School Science* Smithsonian Institution,National Academy of Engineering,National Science Resources Center of the National Academy of Sciences,Institute of Medicine,1998-03-30 With age appropriate inquiry centered curriculum materials and sound teaching practices middle school science can capture the interest and energy of adolescent students and expand their understanding of the world around them *Resources for Teaching Middle School Science* developed by the National Science Resources Center NSRC is a valuable tool for identifying and selecting effective science curriculum materials that will engage students in grades 6 through 8 The volume describes more than 400 curriculum titles that are aligned with the National Science Education Standards This completely new guide follows on the success of *Resources for Teaching Elementary School Science* the first in the NSRC series of annotated guides to hands on inquiry centered curriculum materials and other resources for science teachers The curriculum materials in the new guide are grouped in five chapters by scientific area Physical Science Life Science Environmental Science Earth and Space Science and Multidisciplinary and Applied Science They are also grouped by type core materials supplementary units and science activity books Each annotation of curriculum material includes a recommended grade level a description of the activities involved and of what students can be expected to learn a list of accompanying materials a reading level and ordering information The curriculum materials included in this book were selected by panels of teachers and scientists using evaluation criteria developed for the guide The criteria reflect and incorporate goals and principles of the National Science Education Standards The annotations designate the specific content standards on which these curriculum pieces focus In addition to the curriculum chapters the guide contains six chapters of diverse resources that are directly relevant to middle school science Among these is a chapter on educational software and multimedia programs chapters on books about science and teaching directories and guides to science trade books and periodicals for teachers and students Another section features institutional resources One chapter lists about 600 science centers museums and zoos where teachers can take middle school students for interactive science experiences Another chapter describes nearly 140 professional associations and U S government agencies that offer resources and assistance Authoritative extensive and thoroughly indexed and the only guide of its kind *Resources for Teaching Middle School Science* will be the most used book on the shelf for science teachers school administrators teacher trainers science curriculum specialists advocates of hands on science teaching and concerned parents **Prentice Hall Chemistry** Antony C. Wilbraham,2008 Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science ,2003-11 *Prentice Hall Physical Science Concepts in Action* helps students make the important connection between

the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction *Chemical Flame Inhibition Using Molecular Beam Mass Spectrometry* Alan H. Goode, Ann G. Kim, Arthur W. Hall, Basil Dimitriadis, Carol A. Wilson, Curtis H. Elder, George L. Gates, H. H. Fields, Harry V. Makar, Howard E. Powell, J. R. Allsup, Joan C. Biordi, John C. Franklin, John F. Papp, L. A. Johnson, Martin Hertzberg, Oleg Terichow, Peter R. Tisot, R. H. Sprute, Richard E. Siemens, Charles D. Litton, Charles P. Lazzara, D. J. Kelsh, Edward P. Ferrero, G. V. Sullivan, John H. Perry, L. Y. Sadler, Lee T. Nuzum, Leland C. Marchant, Maurice Deul, Philip C. Good, Ralph D. Fleming, Robert S. Kaplan, W. Hodge Caraway, Waldemar M. Dressel, Adare L. Hill, Cecil Q. Cupps, E. K. Landis, H. J. Coleman, H. J. Lechtenberg, L. Janowski, Robert Lynn Crosby, William Albert Stickney, 1975

Prentice Hall Miller Levine Biology Laboratory Manual a for Students Second Edition 2004 Kenneth Raymond Miller, Joseph S. Levine, Prentice-Hall Staff, 2003-02 Authors Kenneth Miller and Joseph Levine continue to set the standard for clear accessible writing and up to date content that engages student interest Prentice Hall Biology utilizes a student friendly approach that provides a powerful framework for connecting the key concepts a biology Students explore concepts through engaging narrative frequent use of analogies familiar examples and clear and instructional graphics Whether using the text alone or in tandem with exceptional ancillaries and technology teachers can meet the needs of every student at every learning level *Laboratory Guide for Conducting Soil Tests and Plant Analysis Jr.*, J. Benton Jones, 2001-06-28 With the help of this guide you can use obtained test results to evaluate the fertility status of soils and the nutrient element status of plants for crop production purposes It serves as an instructional manual on the techniques used to perform chemical and physical characteristic tests on soils *Laboratory Guide for Conducting Soil Tests and Pl* **Engineering**, 1928 **The Electrical Journal**, 1912 **Laboratory Manual of Physiological Chemistry** Meyer Bodansky, Marion Spencer Fay, 1928

Laboratory Experiments Charles H. Corwin, 1985 **Principles and Practice of Environmental Medicine** A.B. Tarcher, 2013-11-11 Throughout the world scientists and the general with environmental illness Part II presents an over public are concerned about the adverse effects of view of chemical and physical agents commonly toxic agents found in contaminated air water food found in contaminated air water food and soil and soil In the past attention has focused on haz The problem of hazardous wastes is also discussed ards originating in the workplace As a consequence Part III characterizes the body s defense against occupational medicine has become a well recognized such exposure Defenses at the portals of entry are and established clinical discipline Much less atten discussed with emphasis placed on the role of tion has been paid to nonoccupational hazards There nutrition Detoxication and immunologic defense is a growing awareness however of the dangers of mechanisms are described Part IV indicates the exposure to toxic chemical and physical agents in importance of and provides instruction on the the homes community and general environment method of including occupational and

environmen especially for the fetus the infant the very young tal factors in the routine medical history The role of the elderly and the chronically ill those most sus enhanced susceptibility as a factor in an individual s ceptible Environmental medicine fOCUSing on the response to toxic exposure is discussed *Nuclear Science Abstracts* ,1964-04 Chemical Abstracts ,1924 **Good Lighting and the Illuminating Engineer** ,1911 **The Illuminating Engineer** ,1912

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Table of Contents Prentice Hall Flame Test Lab Answers

1. Understanding the eBook Prentice Hall Flame Test Lab Answers
 - The Rise of Digital Reading Prentice Hall Flame Test Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Flame Test Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Flame Test Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Flame Test Lab Answers
 - Personalized Recommendations
 - Prentice Hall Flame Test Lab Answers User Reviews and Ratings
 - Prentice Hall Flame Test Lab Answers and Bestseller Lists

5. Accessing Prentice Hall Flame Test Lab Answers Free and Paid eBooks
 - Prentice Hall Flame Test Lab Answers Public Domain eBooks
 - Prentice Hall Flame Test Lab Answers eBook Subscription Services
 - Prentice Hall Flame Test Lab Answers Budget-Friendly Options
6. Navigating Prentice Hall Flame Test Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Flame Test Lab Answers Compatibility with Devices
 - Prentice Hall Flame Test Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Flame Test Lab Answers
 - Highlighting and Note-Taking Prentice Hall Flame Test Lab Answers
 - Interactive Elements Prentice Hall Flame Test Lab Answers
8. Staying Engaged with Prentice Hall Flame Test Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Flame Test Lab Answers
9. Balancing eBooks and Physical Books Prentice Hall Flame Test Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Flame Test Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Flame Test Lab Answers
 - Setting Reading Goals Prentice Hall Flame Test Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Flame Test Lab Answers
 - Fact-Checking eBook Content of Prentice Hall Flame Test Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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