

Lab#9: Stoichiometry by loss of CO_2

Record the Results Experiment.

Flask#	1	2	3	4	5	6	7
(1). Mass of flask (g)	94.000	94.000	94.000	94.000	94.000	94.000	94.000
(2). V of H_2SO_4 (ml)	20 ml	20 ml	20 ml	20 ml	20 ml	20 ml	20 ml
(3). Mass of flask+ H_2SO_4 (g)	117.934	117.934	117.934	117.934	117.934	117.934	117.934
(4). Mass of H_2SO_4 (g) (3)-(1)	23.934	23.934	23.934	23.934	23.934	23.934	23.934
(5). Mass of beaker (g)	32.000	32.000	32.000	32.000	32.000	32.000	32.000
(6). Mass of beaker+ Na_2CO_3 (g)	33.000g	34.000g	36.000	38.000	40.000g	42.000g	47.000g
(7). Mass of Na_2CO_3 (g) (6)-(5)	1.000g	2.000g	4.000g	6.000g	8.000g	10.000g	12.000g
(8). Mass of reactants (g) (4)+(7)	24.934	25.934	27.934	29.934	31.934	33.934	35.934
(9). Mass of flask+ Products (without CO_2) (g)	113.519	114.103	120.273	121.442	123.293	125.293	127.243
(10). Mass of Products (without CO_2) (g) (9)-(1)	29.519	25.103	26.273	27.442	29.293	31.293	33.243
(11). Mass of CO_2 (g) (8)-(10)	0.415g	.831g	1.661g	2.492	2.641	2.641	2.691
(12). Moles of Na_2CO_3 (7)/(MM of Na_2CO_3)							
(13). Moles of CO_2 (11)/(MM of CO_2)							

Plot a X-Y graph using a graph paper or using Microsoft Excel: Y-axis = moles of CO_2 , X axis = moles of Na_2CO_3 .

Stoichiometry By Loss Of Co₂ Lab Answers

K Morrison



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