

Chapter 9 Stoichiometry Chapter Review Answers

1. Refactor Functions

- Mole ratios are conversion factors that relate the number of moles of one chemical to another chemical involved in a chemical reaction.
- They are formed from the coefficients found in a balanced chemical equation.

2. Mole Ratios as conversion factors

- a. $\frac{2 \text{ mol Cu}}{1 \text{ mol O}_2} \times \frac{1 \text{ mol O}_2}{2 \text{ mol Cu}} = \frac{2 \text{ mol CuO}}{2 \text{ mol Cu}} = \frac{2 \text{ mol Cu}}{2 \text{ mol CuO}} = \frac{1 \text{ mol O}_2}{1 \text{ mol CuO}} = \frac{1 \text{ mol CuO}}{1 \text{ mol O}_2}$. Remember the ratios can be reduced as well!!!

- b. $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol MgF}_2} \cdot \frac{1 \text{ mol MgF}_2}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2}$

- $$3. \quad \frac{105.99 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} \cdot \frac{1 \text{ mol Na}_2\text{CO}_3}{105.99 \text{ g Na}_2\text{CO}_3} \cdot \frac{74.10 \text{ g Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \cdot \frac{1 \text{ mol Ca(OH)}_2}{74.10 \text{ g Ca(OH)}_2} \cdot \frac{40.00 \text{ g NaOH}}{1 \text{ mol NaOH}} \cdot \frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} \cdot \frac{100.09 \text{ g CaCO}_3}{1 \text{ mol CaCO}_3} \cdot \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3}$$

4. Cellular Biology

- Definition: the mass of one mole of a substance
- Molar mass is used as a conversion factor in stoichiometric calculations to convert mass of a substance to moles of the substance or convert moles of a substance to mass of the substance.

5. Hydrogen + Oxygen \rightarrow Water

- a. 5.0 mol H_2 b. 2.5 mol O_2

6. Ethane combustion reaction

- a. 15.8 mol O_2 b. 9.00 mol CO_2 & 13.5 mol H_2O

7. 5.75 g Na & 33.5 g Cl₂

- a. 75.1 mol CO b. 50.1 mol Fe & 75.1 mol CO₂

9. 87.390 g CO or $8.739 \times 10^4 \text{ g CO}$ & 12.610 g H_2 or $1.261 \times 10^4 \text{ g H}_2$

10. Lightning reaction of N_2 & O_2

- a. 1.100 g NO_2 or $1.10 \times 10^3 \text{ g NO}_2$ b. $720. \text{ g NO}$ or $7.20 \times 10^2 \text{ g NO}$

11. CO₂ removal in spacecraft

- a. 42.04 mol NaOH b. 21.02 mol Na_2CO_3 & 21.02 mol H_2O

12. Double replacement AgNO_3 & NaBr

- a. 463 g NaBr b. 845 g AgBr

13. Soda-acid fire extinguisher

- a. 3.058 mol NaHCO_3 b. 3.058 mol CO_2 , 1.529 mol Na_2SO_4 , 3.058 mol H_2O

14. Sulfuric acid-sodium hydroxide neutralization reaction

- $\text{H}_2\text{SO}_4 + 2 \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2 \text{H}_2\text{O}$
- 37 g H_2SO_4
- 53 g Na_2SO_4 & 14 g H_2O

15. Copper & silver nitrate reaction

- 0.0104 mol $\text{Cu}(\text{NO}_3)_2$
- 0.0104 mol Cu & 0.0209 mol AgNO_3

16. Aspirin production

- 13.5 kg $C_3H_8O_3$
- 7.66 kg $C_3H_8O_3$
- 4.29 L $HC_2H_3O_2$

17. Ideal stoichiometry calculations do not account for factors that can affect the relative amounts of product chemicals.

formed in reactions. Ideal stoichiometric calculations are based on the amounts of reactants and products under ideal conditions. Real stoichiometric calculations account for actual conditions under which the reaction occurs.

18. The limiting reactant is the chemical in the reaction in least abundance for the reaction; it will be the chemical reactant that

runs out first in the reaction. The limiting reactant **limits** the amount of chemical product(s) that can be formed. The excess reactant(s) will not be used up completely in the reaction; they will be leftover.

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Arshad Iqbal



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