

Intro to Stoichiometry – Moles to Grams NOTES

Ex. 1: Given 2.5 moles of N_2 , how many grams of NH_3 would be produced?

$$\frac{2.5 \text{ moles } \text{N}_2}{1} \times \frac{2 \text{ moles } \text{NH}_3}{1 \text{ mole } \text{N}_2} \times \frac{17.04 \text{ grams } \text{NH}_3}{1 \text{ mole } \text{NH}_3}$$

Answer: 85.2 grams of NH_3

Ex. 2: Given 4.5 moles of H_2 , how many grams of N_2 will be used?

$$\frac{4.5 \text{ moles } \text{H}_2}{1} \times \frac{1 \text{ moles } \text{N}_2}{3 \text{ mole } \text{H}_2} \times \frac{28.01 \text{ grams } \text{N}_2}{1 \text{ mole } \text{N}_2}$$

Answer: 42.02 grams of N_2

Ex. 3: Given 12.3 grams of NH_3 , how many moles of N_2 were needed?

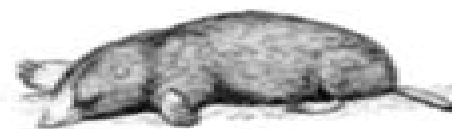
$$\frac{12.3 \text{ grams } \text{NH}_3}{1} \times \frac{1 \text{ mole } \text{NH}_3}{17.04 \text{ grams } \text{NH}_3} \times \frac{1 \text{ mole } \text{N}_2}{2 \text{ moles } \text{NH}_3}$$

Answer: 0.36 moles of N_2

Ex. 4: Given 56 grams of N_2 , how many moles of NH_3 were produced?

$$\frac{56 \text{ grams } \text{N}_2}{1} \times \frac{1 \text{ mole } \text{N}_2}{28.01 \text{ grams } \text{N}_2} \times \frac{2 \text{ moles } \text{NH}_3}{1 \text{ mole } \text{N}_2}$$

Answer: 4.00 moles of NH_3



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Albert A Gayle



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