

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation
- Identify the given (with units) and what you want to find (with units)
- Show set up with units. Check sig figs, give final answer with units and label.

- Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



Before **2 mol** **0 mol** **0 mol**

Change **-2 mol** **+2 mol** **+1 mol**

After **0 mol** **2 mol** **1 mol**

$$36 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.0 \text{ g H}_2\text{O}} = 2.0 \text{ mol H}_2\text{O} \times \frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} = 2 \text{ mol H}_2 \times \frac{2.0 \text{ g H}_2}{1 \text{ mol H}_2} = 4.0 \text{ g H}_2$$

- Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



Before **9.94 mol** **xs mol** **0 mol**

Change **-9.94 mol** **-0.5 mol** **+9.94 mol**

After **0 mol** **0 mol** **9.94 mol**

$$581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

- You eat 180.0 g of glucose (90 M&Ms). If glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to

Stoichiometry Percent Yield Answers

Martha Mackin



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Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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Table of Contents Stoichiometry Percent Yield Answers

1. Understanding the eBook Stoichiometry Percent Yield Answers
 - The Rise of Digital Reading Stoichiometry Percent Yield Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Percent Yield Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Percent Yield Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Percent Yield Answers
 - Personalized Recommendations
 - Stoichiometry Percent Yield Answers User Reviews and Ratings
 - Stoichiometry Percent Yield Answers and Bestseller Lists
5. Accessing Stoichiometry Percent Yield Answers Free and Paid eBooks
 - Stoichiometry Percent Yield Answers Public Domain eBooks
 - Stoichiometry Percent Yield Answers eBook Subscription Services
 - Stoichiometry Percent Yield Answers Budget-Friendly Options

6. Navigating Stoichiometry Percent Yield Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Percent Yield Answers Compatibility with Devices
 - Stoichiometry Percent Yield Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Percent Yield Answers
 - Highlighting and Note-Taking Stoichiometry Percent Yield Answers
 - Interactive Elements Stoichiometry Percent Yield Answers
8. Staying Engaged with Stoichiometry Percent Yield Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Percent Yield Answers
9. Balancing eBooks and Physical Books Stoichiometry Percent Yield Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Percent Yield Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Percent Yield Answers
 - Setting Reading Goals Stoichiometry Percent Yield Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry Percent Yield Answers
 - Fact-Checking eBook Content of Stoichiometry Percent Yield Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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