

## Practice Test Ch 3

## Stoichiometry

Name: \_\_\_\_\_ Per: \_\_\_\_\_

- Remember, this is practice - Do NOT cheat yourself of finding out what you are capable of doing. Be sure you follow the testing conditions outlined below.
- DO NOT USE A CALCULATOR. AP Chem does not allow the use of a calculator for the MC part of the exam, so it is time to start practicing without one. You may use ONLY a periodic table.
- While you should practice working as fast as possible, it is more important at this point in the course, that you practice without a calculator, even if it slows you down. Look for the "easy math" – common factors and rough estimation – do not do "long division" to try to get exact values. Remember it is a MC test, use the answers
- Mark which questions you would like to "go over" when we get to school in September.

1. Balance the following equation:



The balanced equation shows that 1.00 mole of  $\text{NH}_3$  requires \_\_\_\_\_ mole(s) of  $\text{O}_2$ :

- a. 0.57
- b. 1.25
- c. 1.33
- d. 1.75
- e. 3.5

2. Write a balanced equation for the combustion of acetaldehyde,
- $\text{CH}_3\text{CHO}$
- .

When properly balanced, the equation indicates that \_\_\_\_\_ mole(s) of  $\text{O}_2$  are required for each mole of  $\text{CH}_3\text{CHO}$ .

- a. 1
- b. 2
- c. 2.5
- d. 3
- e. 5

3. What is the total mass of products formed when 16 grams of
- $\text{CH}_4$
- is burned with excess oxygen?

- a. 32 g
- b. 36 g
- c. 44 g
- d. 62 g
- e. 80 g

4. Write a balanced equation for the combustion of propane,
- $\text{C}_3\text{H}_8$
- .

When balanced, the equation indicates that \_\_\_\_\_ moles of  $\text{O}_2$  are required for each mole of  $\text{C}_3\text{H}_8$ .

- a. 1.5
- b. 3
- c. 3.5
- d. 5
- e. 8

5. Balance the following equation with the SMALLEST WHOLE NUMBER COEFFICIENTS possible. Select the number that is the
- sum
- of the coefficients in the balanced equation:



- a. 3
- b. 5
- c. 6
- d. 7
- e. 8

6. Calculate the mass of hydrogen formed when 27 g of aluminum reacts with excess hydrochloric acid according to the balanced equation below.



- a. 1.5 g
- b. 2.0 g
- c. 3.0 g
- d. 6.0 g
- e. 12 g

7. How many grams of nitric acid,
- $\text{HNO}_3$
- , can be prepared from the reaction of 128 g of
- $\text{NO}_2$
- with 54.0 g
- $\text{H}_2\text{O}$
- according to the equation below?



- a. 92
- b. 108
- c. 126
- d. 189
- e. 279

8. Which of the following statements is true?

- I. The molar mass of  $\text{CaCO}_3$  is  $100.1 \text{ g mol}^{-1}$ .
- II. 50 g of  $\text{CaCO}_3$  contains  $9 \times 10^{23}$  oxygen atoms.
- III. A 200 g sample of  $\text{CaCO}_3$  contains 2 moles of  $\text{CaCO}_3$ .

- a. I only
- b. II only
- c. III only
- d. I and III only
- e. I, II, and III

# Stoichiometry Practice Test With Answers

**Kathy A. Zahler**



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