

Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



Reactant side

4 Fe atoms

4 S atoms

7(2) = 14 O atoms

Product side

4 Fe atoms

4 S atoms

2(3) + 4(2) = 14 O atoms

b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

m FeS = 2.0 g

Req'd. = m O₂

Solution:

molar mass of FeS = 88 g/mol

mole ratio = 4 mol FeS: 7 mol O₂

molar mass of O₂ = 32 g/mol

$$m_{O_2} = 2.0 \cancel{g_{FeS}} \left(\frac{1 \cancel{mol_{FeS}}}{88 \cancel{g_{FeS}}} \right) \left(\frac{7 \cancel{mol_{O_2}}}{4 \cancel{mol_{FeS}}} \right) \left(\frac{32 \cancel{g_{O_2}}}{1 \cancel{mol_{O_2}}} \right) = 1.3 g_{O_2}$$

2. Rusting involves the reaction of iron, Fe with oxygen, O₂ to form iron (III) oxide, Fe₂O₃. (20pts)

a) Write the balanced chemical equation.



Reactant side

4 Fe atoms

3(2) = 6 O atoms

Product side

4 Fe atoms

2(3) = 6 O atoms

b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

4 mol Fe: 3 mol O₂

Product to each reactant:

2 mol Fe₂O₃: 4 mol Fe

Stoichiometry Problems Packet Answers

Steven S. Zumdahl, Paul B. Kelter



Stoichiometry Problems Packet Answers:

Solving Problems in Chemistry Rod O'Connor, Charles Mickey, Alton Hassell, 1977 **Water Quality Engineering** Mark M. Benjamin, Desmond F. Lawler, 2013-07-01 Explains the fundamental theory and mathematics of water and wastewater treatment processes By carefully explaining both the underlying theory and the underlying mathematics this text enables readers to fully grasp the fundamentals of physical and chemical treatment processes for water and wastewater Throughout the book the authors use detailed examples to illustrate real world challenges and their solutions including step by step mathematical calculations Each chapter ends with a set of problems that enable readers to put their knowledge into practice by developing and analyzing complex processes for the removal of soluble and particulate materials in order to ensure the safety of our water supplies Designed to give readers a deep understanding of how water treatment processes actually work Water Quality Engineering explores Application of mass balances in continuous flow systems enabling readers to understand and predict changes in water quality Processes for removing soluble contaminants from water including treatment of municipal and industrial wastes Processes for removing particulate materials from water Membrane processes to remove both soluble and particulate materials Following the discussion of mass balances in continuous flow systems in the first part of the book the authors explain and analyze water treatment processes in subsequent chapters by setting forth the relevant mass balance for the process reactor geometry and flow pattern under consideration With its many examples and problem sets Water Quality Engineering is recommended as a textbook for graduate courses in physical and chemical treatment processes for water and wastewater By drawing together the most recent research findings and industry practices this text is also recommended for professional environmental engineers in search of a contemporary perspective on water and wastewater treatment processes Subject Guide to Forthcoming Books, 1983 Presents by subject the same titles that are listed by author and title in Forthcoming books **Biochemistry of Cell Walls and Membranes** C. Fred Fox, Kenneth Burton, David Chilton Phillips, Sir Hans Kornberg, 1975 **Government Reports Annual Index**, 1983 **Transactions** Biochemical Society (Great Britain), 2005 Sassy Stoichiometry Problems Julie C. Gilbert, 2021-03-14 Need more Stoichiometry practice Stoichiometry has been striking fear into the hearts of chemistry students for ages The best way to conquer something is to practice it Inside you ll find Brief descriptions of each type of ideal stoichiometry and limiting reactant stoichiometry 4 ideal stoichiometry worksheets broken down by type with keys and explanations 4 ideal stoichiometry self quizzes with their answer keys 2 limiting reactant stoichiometry worksheets with keys and explanations 2 limiting reactant stoichiometry self quizzes with answer keys 2 mixed stoichiometry self tests with answer keys This is a companion workbook for the 5 Steps to Surviving Chemistry book However you do not need to have read that book to find this workbook useful **Stoichiometry Problems**, 2016 You extend your study of stoichiometry to consider more complex problems involving volume molecules and energy Stoichiometry Michell J. Sienko, 1964 A Stoichiometry Unit David

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Problem Solving Strategies in Stoichiometry Used by Students with Different Learning Approaches and Different Conceptual Understanding Levels Hala Raji Barakat,1998 Most of the strategies identified were either algorithmic correct strategies or incorrect strategies with incorrect answers where conceptual misunderstanding impeded successful problem solving The more conceptual thinkers could use the concepts to solve problems successfully in the test and argue correctly in the interview Moreover the learning approach was found not to be related either to conceptual understanding or to problem solving strategies A major educational implication was that instruction should give more emphasis to conceptual

understanding to enhance meaningful problem solving Also exams should include both the conceptual and the algorithmic aspects of problem solving *Solving Problems in Chemistry* R. O'Connor, C. Mickey, 1974

The book delves into Stoichiometry Problems Packet Answers. Stoichiometry Problems Packet Answers is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Stoichiometry Problems Packet Answers, encompassing both the fundamentals and more intricate discussions.

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