

## Stoichiometry Worksheet #2

1. If 1.87g of aluminum is reacted with an excess of copper(II) sulfate and 3.74g of copper metal are actually collected, what is the percent yield?

$$2\text{Al}_{(s)} + 3\text{CuSO}_{4(aq)} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}_{(s)}$$

$$1.87\text{g Al} \left( \frac{1 \text{ mol Al}}{26.98\text{g Al}} \right) \left( \frac{3 \text{ mol Cu}}{2 \text{ mol Al}} \right) \left( \frac{63.55\text{g Cu}}{1 \text{ mol Cu}} \right) = \boxed{6.61\text{g Cu}} \text{ Theoretical Yield}$$

$$\frac{3.74\text{g Cu (Actual Yield)}}{6.61 \text{ (Theoretical)}} \times 100 = \boxed{56.6\% \text{ Yield}}$$

2. 0.531 mol of iron(III) oxide are reacted with an excess of carbon monoxide and 57.8g of iron is produced. What is the percent yield? Hint: The other product of the reaction is carbon dioxide.

$$\text{Fe}_2\text{O}_{3(s)} + 3\text{CO}_{(g)} \rightarrow 3\text{CO}_{2(g)} + 2\text{Fe}_{(s)}$$

$$.531 \text{ mol Fe}_2\text{O}_3 \left( \frac{2 \text{ mol Fe}}{1 \text{ mol Fe}_2\text{O}_3} \right) \left( \frac{55.85\text{g Fe}}{1 \text{ mol Fe}} \right) = \boxed{59.3\text{g Fe}} \text{ Theoretical Yield}$$

$$\frac{57.8\text{g (Actual)}}{59.3\text{g (Theor)}} \times 100 = \boxed{97.5\% \text{ Yield}}$$

3. Calcium carbonate can be decomposed by heating to produce solid calcium oxide and carbon dioxide gas. If you wanted to produce 5.55L of carbon dioxide gas, what mass of calcium carbonate would have to be used? Assume 100% yield and conditions at STP.

$$\text{CaCO}_3 \xrightarrow{\Delta} \text{CaO}_{(s)} + \text{CO}_{2(g)}$$

$$5.55 \text{ L CO}_2 \left( \frac{1 \text{ mol CO}_2}{22.4 \text{ L CO}_2} \right) \left( \frac{1 \text{ mol CaCO}_3}{1 \text{ mol CO}_2} \right) \left( \frac{100.09\text{g CaCO}_3}{1 \text{ mol CaCO}_3} \right) = \boxed{24.8\text{g CaCO}_3}$$

4. Aluminum chloride, an inexpensive reagent used in many industrial processes, is made by reacting scrap aluminum with chlorine. If 2.70g of aluminum are reacted with 4.05g of chlorine, which is the limiting reagent and how many grams of the excess reagent remain when the reaction is complete?

$$2\text{Al}_{(s)} + 3\text{Cl}_{2(g)} \rightarrow 2\text{AlCl}_{3(s)}$$

$$2.70\text{g Al} \left( \frac{1 \text{ mol Al}}{26.98\text{g Al}} \right) \left( \frac{2 \text{ mol AlCl}_3}{2 \text{ mol Al}} \right) = .100 \text{ mol AlCl}_3$$

$$4.05\text{g Cl}_2 \left( \frac{1 \text{ mol Cl}_2}{70.90\text{g Cl}_2} \right) \left( \frac{2 \text{ mol AlCl}_3}{3 \text{ mol Cl}_2} \right) = .0381 \text{ mol AlCl}_3$$

Aluminum Produced Cl<sub>2</sub> is limiting

$$.100 \text{ mol Al} - .0381 \text{ mol Al} = .0619 \text{ mol Al (Produced)}$$

$$.0619 \text{ mol AlCl}_3 \left( \frac{2 \text{ mol Al}}{2 \text{ mol AlCl}_3} \right) \left( \frac{26.98\text{g Al}}{1 \text{ mol Al}} \right) = 1.67\text{g Al in Excess}$$

# Stoichiometry Quiz 2 Answer Key Chemistry

**Mike Jess**



## **Stoichiometry Quiz 2 Answer Key Chemistry:**

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*General Chemistry* B. Richard Siebring, Mary Ellen Schaff, 1980

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The Periodic Table More About the Periodic Table Some Important Trends VIII Solids Liquids and Gases Gases Intermolecular Forces Phase Changes Energy and Phase Changes IX Solutions Solutions Concentrations Solubility and Saturation X Kinetics and Equilibrium Kinetics Factors that Affect Reaction Rate Reversible Reactions and Chemical Equilibrium Le Chatelier's Principle XI Acids and Bases Acids and Bases Titration XII Redox and Electrochemistry Oxidation and Reduction Electrochemistry XIII Organic Chemistry Hydrocarbons Functional Groups XIV Laboratory Safety Rules Accuracy Significant Figures Lab Procedures Laboratory Equipment XV Practice Tests     *The Best Test Preparation for the College Board Achievement Test in Chemistry* Research and Education Association, 1987-02-20 Master the SAT II Chemistry Subject Test and score higher Our test experts show you the right way to prepare for this important college exam REA's SAT II Chemistry test prep covers all chemistry topics to appear on the actual exam including in depth coverage of the laws of chemistry properties of solids gases and liquids chemical reactions and more The book features 6 full length practice SAT II Chemistry exams Each practice exam question is fully explained to help you better understand the subject material Use the book's Periodic Table of Elements for speedy look up of the properties of each element Follow up your study with REA's proven test taking strategies powerhouse drills and study schedule that get you ready for test day DETAILS Comprehensive review of every chemistry topic to appear on the SAT II subject test Flexible study schedule tailored to your needs Packed with proven test tips strategies and advice to help you master the test 6 full length practice SAT II Chemistry Subject tests Each test question is answered in complete detail with easy to follow easy to grasp explanations The book's handy Periodic Table of Elements allows for quick answers on the elements appearing on the exam TABLE OF CONTENTS About Research and Education Association Independent Study Schedule CHAPTER 1 ABOUT THE SAT II CHEMISTRY SUBJECT TEST About This Book About The Test How To Use This Book Format of the SAT II Chemistry Scoring the SAT II Chemistry Score Conversion Table Studying for the SAT II Chemistry Test Taking Tips CHAPTER 2 COURSE REVIEW Gases Gas Laws Gas Mixtures and Other Physical Properties of Gases Dalton's Law of Partial Pressures Avogadro's Law The Mole Concept Avogadro's Hypothesis Chemical Compounds and Formulas Mole Concept Molecular Weight and Formula Weight Equivalent Weight Chemical Composition Stoichiometry Weight and Volume Calculations Balancing Chemical Equations Calculations Based on Chemical Equations Limiting Reactant Calculations Solids Phase Diagram Phase Equilibrium Properties of Liquids Density Colligative Properties of Solutions Raoult's Law and Vapor Pressure Osmotic Pressure Solution Chemistry Concentration Units Equilibrium The Law of Mass Action Kinetics and Equilibrium Le Chatelier's Principle and Chemical Equilibrium Acid Base Equilibria Definitions of Acids and Bases Ionization of Water pH Dissociation of Weak Electrolytes Dissociation of Polyprotic Acids Buffers Hydrolysis Thermodynamics I Bond Energies Some Commonly Used Terms in Thermodynamics The First Law of Thermodynamics Enthalpy Hess's Law of Heat Summation Standard States Heat of Vaporization and Heat of Fusion Thermodynamics II Entropy The Second Law of Thermodynamics Standard Entropies and

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know when taking the exam Be sure to take the remaining practice tests to further test yourself and become familiar with the format of the SAT II Chemistry Subject Test

### When Should I Start Studying

It is never too early to start studying for the SAT II Chemistry test The earlier you begin the more time you will have to sharpen your skills Do not procrastinate Cramming is not an effective way to study since it does not allow you the time needed to learn the test material The sooner you learn the format of the exam the more comfortable you will be when you take the exam

### FORMAT OF THE SAT II CHEMISTRY

The SAT II Chemistry is a one hour exam consisting of 85 multiple choice questions The first part of the exam consists of classification questions This question type presents a list of statements or questions that you must match up with a group of choices lettered A through E Each choice may be used once more than once or not at all The exam then shifts to relationship analysis questions which you will answer in a specially numbered section of your answer sheet You will have to determine if each of two statements is true or false and if the second statement is a correct explanation of the first The last section is composed strictly of multiple choice questions with choices lettered A through E

### Material Tested

The following chart summarizes the distribution of topics covered on the SAT II Chemistry Subject Test

Topic	Percentage	Number of Questions
Atomic	however do not review too much at any one time	
Concentrate on one problem area at a time by reviewing the questions and explanations and by studying our review until you are confident you completely understand the material		
Keep track of your scores By doing so you will be able to gauge your progress and discover general weaknesses in particular sections You should carefully study the reviews that cover your areas of difficulty as this will build your skills in those areas		

### TEST TAKING TIPS

Although you may be unfamiliar with standardized tests such as the SAT II Chemistry Subject Test there are many ways to acquaint yourself with this type of examination and help alleviate your test taking anxieties Become comfortable with the format of the exam When you are practicing to take the SAT II Chemistry Subject Test simulate the conditions under which you will be taking the actual test Stay calm and pace yourself After simulating the test only a couple of times you will boost your chances of doing well and you will be able to sit down for the actual exam with much more confidence Know the directions and format for each section of the test Familiarizing yourself with the directions and format of the exam will not only save you time but will also ensure that you are familiar enough with the SAT II Chemistry Subject Test to avoid nervousness and the mistakes caused by being nervous Do your scratchwork in the margins of the test booklet You will not be given scrap paper during the exam and you may not perform scratchwork on your answer sheet Space is provided in your test booklet to do any necessary work or draw diagrams If you are unsure of an answer guess However if you do guess guess wisely Use the process of elimination by going through each answer to a question and ruling out as many of the answer choices as possible By eliminating three answer choices you give yourself a fifty fifty chance of answering correctly since there will only be two choices left from which to make your guess Mark your answers in the appropriate spaces on the answer sheet Fill in the oval that corresponds to your answer darkly completely and neatly You can change your answer but

remember to completely erase your old answer Any stray lines or unnecessary marks may cause the machine to score your answer incorrectly When you have finished working on a section you may want to go back and check to make sure your answers correspond to the correct questions Marking one answer in the wrong space will throw off the rest of your test whether it is graded by machine or by hand You don t have to answer every question You are not penalized if you do not answer every question The only penalty results from answering a question incorrectly Try to use the guessing strategy but if you are truly stumped by a question remember that you do not have to answer it Work quickly and steadily You have a limited amount of time to work on each section so you need to work quickly and steadily Avoid focusing on one problem for too long Before the Test Make sure you know where your test center is well in advance of your test day so you do not get lost on the day of the test On the night before the test gather together the materials you will need the next day Your admission ticket Two forms of identification e g driver s license student identification card or current alien registration card Two No 2 pencils with erasers Directions to the test center A watch if you wish but not one that makes noise as it may disturb other test takers On the day of the test you should wake up early after a good night s rest and have breakfast Dress comfortably so that you are not distracted by being too hot or too cold while taking the test Also plan to arrive at the test center early This will allow you to collect your thoughts and relax before the test and will also spare you the stress of being late If you arrive after the test begins you will not be admitted to the test center and you will not receive a refund During the Test When you arrive at the test center try to find a seat where you feel most comfortable Follow all the rules and instructions given by the test supervisor If you do not you risk being dismissed from the test and having your scores canceled Once all the test materials are passed out the test instructor will give you directions for filling out your answer sheet Fill this sheet out carefully since this information will appear on your score report After the Test When you have completed the SAT II Chemistry Subject Test you may hand in your test materials and leave Then go home and relax When Will I Receive My Score Report and What Will It Look Like You should receive your score report about five weeks after you take the test This report will include your scores percentile ranks and interpretive information

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In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Stoichiometry Quiz 2 Answer Key Chemistry**," a mesmerizing literary creation penned by a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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### Table of Contents **Stoichiometry Quiz 2 Answer Key Chemistry**

1. Understanding the eBook **Stoichiometry Quiz 2 Answer Key Chemistry**
  - The Rise of Digital Reading **Stoichiometry Quiz 2 Answer Key Chemistry**
  - Advantages of eBooks Over Traditional Books
2. Identifying **Stoichiometry Quiz 2 Answer Key Chemistry**
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an **Stoichiometry Quiz 2 Answer Key Chemistry**
  - User-Friendly Interface
4. Exploring eBook Recommendations from **Stoichiometry Quiz 2 Answer Key Chemistry**
  - Personalized Recommendations
  - **Stoichiometry Quiz 2 Answer Key Chemistry** User Reviews and Ratings
  - **Stoichiometry Quiz 2 Answer Key Chemistry** and Bestseller Lists

5. Accessing Stoichiometry Quiz 2 Answer Key Chemistry Free and Paid eBooks
  - Stoichiometry Quiz 2 Answer Key Chemistry Public Domain eBooks
  - Stoichiometry Quiz 2 Answer Key Chemistry eBook Subscription Services
  - Stoichiometry Quiz 2 Answer Key Chemistry Budget-Friendly Options
6. Navigating Stoichiometry Quiz 2 Answer Key Chemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Stoichiometry Quiz 2 Answer Key Chemistry Compatibility with Devices
  - Stoichiometry Quiz 2 Answer Key Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Stoichiometry Quiz 2 Answer Key Chemistry
  - Highlighting and Note-Taking Stoichiometry Quiz 2 Answer Key Chemistry
  - Interactive Elements Stoichiometry Quiz 2 Answer Key Chemistry
8. Staying Engaged with Stoichiometry Quiz 2 Answer Key Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Stoichiometry Quiz 2 Answer Key Chemistry
9. Balancing eBooks and Physical Books Stoichiometry Quiz 2 Answer Key Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Stoichiometry Quiz 2 Answer Key Chemistry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Quiz 2 Answer Key Chemistry
  - Setting Reading Goals Stoichiometry Quiz 2 Answer Key Chemistry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry Quiz 2 Answer Key Chemistry
  - Fact-Checking eBook Content of Stoichiometry Quiz 2 Answer Key Chemistry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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