

Quiz 6: Stoichiometry SCH3U Name:

 T C A **Assessment of Learning (Product)**
13 4 6

1. The following questions pertain to the following chemical equation:



- a) If 3.94×10^{24} molecules of HNO_3 are mixed with 54.2 g of sulfur monoxide what is the limiting reactant? (show all work) (6T, 1C)
b) What is the theoretical yield of H_2SO_4 in molecules, expected in part a? (1T, 1C)
c) What mass of NO_2 is expected in part a? (3T, 1C)
d) What mass of the excess reactant is left over after the reaction? (3T, 1C)

2. During a lab a student reacts 6.51 g of $\text{C}_2\text{H}_6\text{O}$ with excess oxygen.

a. Write a balanced complete combustion reaction (2A)

b. Calculate the actual yield of CO_2 in grams, that forms, if the reaction proceeds with a 54.8% yield (4A)

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