

Ch 9 stoichiometry Guided Reading Notes

9.1

Objectives

- Define Stoichiometry
- Describe the importance of the mole ratio in stoichiometric calculations
- Write a mole ratio relating two substances in a chemical reaction.

In stoichiometry problems you will be given the amount of at least one of the substances in a chemical equation and asked to find the amount of one or more substances in the equation

The given amount may be expressed in _____ or _____

The Unknown amount may be expressed in _____ or _____

What are the 4 types of stoichiometry problems?

What conversion factor do you use to change grams to moles?

All stoichiometry problems involve the use of the _____

Where do you get the mole ratio from?

What does the mole ratio enable you to do?

Can you perform a stoichiometry problem with an unbalanced chemical equation? Explain

9.1 assignment p 285 1 to 4 p 305 1 to 3 ws pages 1,2

9.2

Objectives

- Calculate the amount in moles of a reactant or product from the amount in moles of a different reactant or product
- Calculate the mass of a reactant or product from the amount in moles of a different reactant or product
- Calculate the amount in moles of a reactant or product from the mass of a different reactant or product
- Calculate the mass of a reactant or product from the mass of a different mass or product

Given moles, asked to find moles, use mole ratio

Do p 289, 1,2 Here are the balanced equations $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$, $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$

Given mole, asked to find grams, Use molar mass to convert grams to moles, then use mole Ratio

Do p 291, 1 and 2

Given grams, asked to find moles, use molar mass to convert grams to moles, then mole ratio

Do p 293, 1,2 Here is the balanced equation $2\text{HgO} \rightarrow 2\text{Hg} + \text{O}_2$

Stoichiometry Section 9 1 Guided Reading Answers

Iris Stovall, Donald J. DeCoste



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