

# Stoichiometry Study Guide

Balance the following equations. Then use the balanced equation to answer questions #1-5.

1) 10 moles of  $\text{H}_2\text{O}$  are

a) 180 grams

b) 360 grams

c) 540 grams

2) 10 moles of  $\text{H}_2\text{O}$  are

a) 180 grams

b) 360 grams

c) 540 grams

Balance the following equation.

3) 10 moles of  $\text{H}_2\text{O}$  are

a) 180 grams

b) 360 grams

c) 540 grams

4) 10 moles of  $\text{H}_2\text{O}$  are

a) 180 grams

b) 360 grams

c) 540 grams

Name \_\_\_\_\_ Period \_\_\_\_\_ Date \_\_\_\_\_

## Stoichiometry Study Guide

Fill in the blanks below.

1 mole = \_\_\_\_\_ grams.

Balance the following equation. Then use the balanced equation to answer the questions #1-5.



1) Find the formula for the following compounds:

a) NaCl to  $\text{O}_2$

b) NaCl to  $\text{NaClO}_2$

c)  $\text{O}_2$  to  $\text{NaClO}_2$

2) If the above chemical reaction is done as an experiment using 4.4 moles of each reactant, calculate how much  $\text{NaClO}_2$  will be produced?

a) Given 4.4 moles of NaCl, how many moles of  $\text{NaClO}_2$  will be produced?

b) Given 4.4 moles of  $\text{O}_2$ , how many moles of  $\text{NaClO}_2$  will be produced?

c) Define limiting reactant & excess reactant.

d) Which reactant is the limiting reactant? Which reactant is the excess reactant?

3) If the above chemical reaction is done as an experiment using 10 moles of  $\text{O}_2$ , calculate the percent yield given that the actual yield of  $\text{NaClO}_2$  is 8.8 moles.

a) What is the theoretical yield in grams of  $\text{NaClO}_2$  given 10 moles of the  $\text{O}_2$ ?

b) Write the formula for Percent Yield.

c) What is the percent yield given that the actual yield of  $\text{NaClO}_2$  is 8.8 moles?

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**Judd E. Hollander**



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