

# Stoichiometry Worksheet and Key

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1. How many moles of  $\text{O}_2$  will be formed from 1.65 moles of  $\text{KClO}_3$ ?

1.65 mol $\text{KClO}_3$	_____ mol $\text{O}_2$	= _____ mol $\text{O}_2$
	_____ mol $\text{KClO}_3$	

2. How many moles of  $\text{KClO}_3$  are needed to make 3.50 moles of  $\text{KCl}$ ?

3.50 mol $\text{KCl}$		= _____ mol $\text{KClO}_3$

3. How many moles of  $\text{KCl}$  will be formed from 2.73 moles of  $\text{KClO}_3$ ?

_____		= _____



4. How many moles of  $\text{Fe}_2\text{O}_3$  are produced when 0.275 moles of  $\text{Fe}$  is reacted?

0.275 mol $\text{Fe}$		= _____ mol $\text{Fe}_2\text{O}_3$

5. How many moles of  $\text{Fe}_2\text{O}_3$  are produced when 31.0 moles of  $\text{O}_2$  is reacted?

_____		= _____

6. How many moles of  $\text{O}_2$  are needed to react with 8.9 moles of  $\text{Fe}$ ?

_____		= _____



7. How many moles of  $\text{O}_2$  are produced when 1.26 moles of  $\text{H}_2\text{O}$  is reacted?

8. How many moles of  $\text{H}_2\text{O}$  are needed to produce 55.7 moles of  $\text{H}_2$ ?

9. If enough  $\text{H}_2\text{O}$  is reacted to produce 3.40 moles of  $\text{H}_2$ , then how many moles of  $\text{O}_2$  must have been made?  
(a bit challenging, but just think about it and you can probably figure it out)

# Stoichiometry Test B Answers 1

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