

# Stoichiometry: Assessment Review

Name \_\_\_\_\_

**Skill** – Student can use a balanced equation to determine mole to mole ratios

- 1) Given this equation:  $N_2 + 3 H_2 \rightarrow 2 NH_3$ , write the following molar ratios:
  - a.  $N_2 : H_2$
  - b.  $N_2 : NH_3$
  - c.  $H_2 : NH_3$
- 2) Given the following equation:  $8 H_2 + S_8 \rightarrow 8 H_2S$ , write the following molar ratios:
  - a.  $H_2 : H_2S$
  - b.  $H_2 : S_8$
  - c.  $H_2S : S_8$
- 3) Answer the following questions for this equation:  $2 H_2 + O_2 \rightarrow 2 H_2O$ 
  - a. What is the  $H_2 : H_2O$  molar ratio?
  - b. With 20 moles of  $H_2$  on hand & plenty of  $O_2$ , how many moles of  $H_2O$  could you make?
  - c. What is the  $O_2 : H_2O$  molar ratio?
  - d. With 20 moles of  $O_2$  and enough  $H_2$ , how many moles of  $H_2O$  could you make?
- 4) Answer the following questions for this equation:  $N_2 + 3 H_2 \rightarrow 2 NH_3$ 
  - a. If you used 1 mole of  $N_2$ , how many moles of  $NH_3$  could be produced?
  - b. If 10 moles of  $NH_3$  were produced, how many moles of  $N_2$  would be required?
  - c. If 3.00 moles of  $H_2$  were used, how many moles of  $NH_3$  would be made?
  - d. If 0.600 moles of  $NH_3$  were produced, how many moles of  $H_2$  are required?

**Skill** – Student can use a balanced equation, given the grams of reactant, to calculate the grams of another substance

- 5) Magnesium reacts with hydrochloric acid according to the following equation:



- If 72.92 grams of hydrochloric acid react with excess magnesium, how many moles of hydrogen gas will be produced?
- 6) Aluminum reacts with hydrobromic acid to produce aluminum bromide and hydrogen gas. Write a balanced equation for the reaction and calculate the number of moles of hydrobromic acid required to react with 29.47 grams of aluminum.

# Stoichiometry Vocabulary Review Answer Key Work Sheet

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