

# Unit: Stoichiometry WS 6 - "Test Review"



Sample questions/Important Topics: Use dimensional analysis to perform the following calculations. Show all of your work.

1. How many moles of calcium nitrate would react with 4.55 moles of Chromium (III) sulfate to produce calcium sulfate and chromium(III) nitrate?



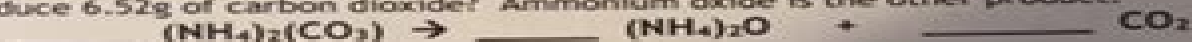
2. How many moles of Zinc would be produced from the reaction of 13.2 moles of copper with excess zinc acetate? Copper(I) acetate is the other product.



3. What mass of Gallium chloride would react with excess fluorine to produce 78.9 g of chlorine? Gallium fluoride is the other product.



4. How many grams of ammonium carbonate are needed to decompose in order to produce 6.52g of carbon dioxide? Ammonium oxide is the other product.



5. What volume of nitrogen gas at STP would react with 37.2 g of magnesium to produce magnesium nitride?



6. How many molecules of sulfur trioxide would react with 68.9 L of water at STP to produce sulfuric acid?



7. What is the maximum number of grams of silver phosphate that can be produced from the reaction of 287 g of silver chlorate with 52.8 g of sodium phosphate?



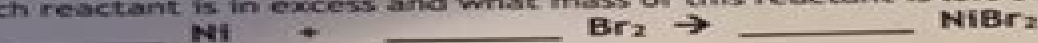
8. How many formula units of iron(III) oxide can be produced from the reaction of 108 g of iron with 72.1 L of oxygen gas at STP?



9. How many moles of barium hydroxide can be produced from the reaction of 4.234 moles of barium with 8.902 moles of water according to the following reaction?



10. a) What is the maximum number of grams of nickel bromide that can be produced from the reaction of 67.8 g of nickel with 37.3 g of bromine?  
b) Which reactant is in excess and what mass of this reactant is left over?



# Stoichiometry Ws 6 Review Answers

**Rachel Sandford**



## **Stoichiometry Ws 6 Review Answers:**

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web 1 how does the comic strip define a polar molecule 2 define electronegativity as you understand it after reading the first two pages of the comic strip 3 what is the artist trying to represent by two polar bears arm wrestling or two penguins arm

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



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