Study Guide - Chapter 3: Stoichiometry

Test Date: Thursday, 12/18/14

- Chemical Equations
 - Writing word equations
 - remember to balance & include the states of matter
 - Balancing chemical equations w/coefficients
 - 5 types of reactions
 - Combination/synthesis, Decomposition, Combustion, Single Replacement, Double Replacement
 - identify what type of reaction is happening from a given equation
 - predict the products when given the reactants or told the type of reaction...
 - C_xH_y + O₂ (g) combustion forms products of CO₂(g) and H₂O (g)
 - Synthesis of metal + nonmetal forms an ionic compound (criss corss charges for formula)
 - Decomposition breaks compound itto its elements (watch out for diatomic BrINCIHOFs or P₄ / S₈ solids)
- Formula Weight (unit: amu) = Molar Mass (unit: g/mol)
 - Calculations of formula weights and/or molar masses.
 - % by mass = % composition calculations
- Determining Empirical & Molecular Formulas
 - Know the difference between an empirical and a molecular formula
 - Three types of problems
 - Normal problems will give % composition data or mass data to start
 - Combustion Analysis (like the quiz)
 - Hydrate Determination
 - Use the rhyme to help you remember the steps!
- Stoichiometric Calculations
 - Dimensional Analysis!!!
 - Avogadro's number (use for converting to or from # of atoms/molecules) → after molecules you can go to specific atom within the compound by looking at the chemical formula (ex: 4 H atoms per 1 molec. CH₄)
 - use the molar mass to go from mass to moles and/or moles to mass
 - Molar Ratios: use coefficients from a balanced chemical equation to go between any two different substances in a reaction (section 3.6)
- Limiting Reactant Problems
 - through calculations, identify which is the limiting reactant and which is the excess reactant
 - calculate the amount of product formed by the limiting reactant (this is the maximum or theoretical yield but you ALWAYS choose the LESSER quantity calculated)
 - determine how much of the excess reactant is used and how much is left over after the reaction is completed
 - calculate % yield
 - calculate theoretical yield and given the actual yield (from lab) of the substance
 - given % yield, you can determine actual yield collected in the lab

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