#### Stoichiometry Worksheet # 1

 Glucose is used as a source of energy by the human body. The overall reaction in the body is C6H12O6+6O2 → 6CO2 + 6H2O

Calculate the number of moles of oxygen needed to oxidize 12.5 g of C6H12O6.

Ammonia is synthesized from hydrogen and nitrogen according to the following equation.

$$N_2 + 3H_2 \rightarrow 2NH_3$$

If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH<sub>3</sub>) can be produced?

3) Assume that in the decomposition of potassium chlorate, KCLO<sub>3</sub>,

80.5 moles of O: form. How many grams of potassium chlorite, the other product would be formed?

4) In a single displacement reaction,

9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

5) The compound "Cisplatin" PtCl<sub>2</sub>(NH<sub>2</sub>)<sub>2</sub>, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction

$$K_2PtCl_4 + 2NH_3 \longrightarrow 2KCl + PtCl_2(NH_3)_2$$

- A. How many moles of "cisplatin" can be produced from 2.50 g K2PtCl4?
- B. How many moles of NH3 would be needed if 3.8 moles of K2PtCl4 were present?

## **Stoichiometry Chapter 12 Key**

Richard H. Langley, John Moore

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Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives. These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance **Biochemistry** John L. Tymoczko, Jeremy M. Berg, Lubert Stryer, 2010 Derived from the classic text originated by Lubert Stryer and continued by John Tymoczko and Jeremy Berg Biochemistry A Short Course offers that bestseller s signature writing style and physiological emphasis while focusing on the major topics taught in a one semester biochemistry Thermodynamics Stephen R. Turns, Laura L. Pauley, 2020-02-27 Fully revised to match the more traditional course sequence of course materials this full color second edition presents the basic principles and methods of thermodynamics using a clear and engaging style and a wealth of end of chapter problems It includes five new chapters on topics such as mixtures psychometry chemical equilibrium and combustion and discussion of the Second Law of Thermodynamics has been expanded and divided into two chapters allowing instructors to introduce the topic using either the cycle analysis in Chapter 6 or the definition of entropy in Chapter 7 Online ancillaries including new LMS testbanks a password protected solutions manual prepared PowerPoint lecture slides instructional videos and figures in electronic format are available at www cambridge org thermo Foundations of College Chemistry, Alternate Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis Teaching Science for Understanding James Joseph Gallagher, 2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities **Essentials of Chemical Reaction Engineering** H. Scott Fogler, 2011 Accompanying DVD ROM contains many realistic interactive simulations Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change Chemistry: Core Concepts, 3rd Edition Allan Blackman, Daniel

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