Stoichiometry

Solve the following problems:

- Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)
- a) Translate the problem to chemical equation.

$$4FeS + 7O_2 \longrightarrow 2Fe_2O_2 + 4SO_2$$

Reactant side	Product side
4 Fe atoms	4 Fe atoms
4 S atoms	4 S atoms
7(2) =14 O atoms	2(3) +4(2) =14 O atoms

b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given: m FeS = 2.0 gReq'd. = $m O_2$

Solution: molar mass of FeS = 88 g/molmole ratio = 4 mol FeS: 7 mol O_2 molar mass of $O_2 = 32 \text{ g/mol}$

$$m_{O_2} = 2.0 g_{FeS} \ \left(\frac{1 \, \text{mot}_{Fes}}{88 g_{FeS}} \right) \left(\frac{7 \, \text{mot}_{O_2}}{4 \, \text{mot}_{FeS}} \right) \left(\frac{32 g_{O_2}}{1 \, \text{mot}_{O_2}} \right) = 1.3 g_{O_2}$$

- Rusting involves the reaction of iron, Fe with oxygen, O₂ to form iron (III) oxide, Fe₂O₃. (20pts)
 - a) Write the balanced chemical equation.

 Reactant side
 Product side

 4 Fe atoms
 4 Fe atoms

 3(2) =6 O atoms
 2(3) =6 O atoms

b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

Product to each reactant:

4 mol Fe: 3 mol O_2

2 mol FeO₂:4 mol Fe

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