

## Unit 6: Stoichiometry Homework Packet

Name: \_\_\_\_\_

Period: \_\_\_\_\_

### Mole Ratios:

1. What is a mole ratio?

\_\_\_\_\_

2. The coefficients in a chemical equation indicate the relationship among \_\_\_\_\_ of reactants and products.

3. Write 3 possible mole ratios from this equation:  $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$

4. Write 4 possible mole ratios from this equation:  $\text{C}_2\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

5. Write a balanced chemical equation when nitrogen reacts with hydrogen to produce nitrogen trihydride. Determine 3 possible mole ratios.

6. Stoichiometry is based on the \_\_\_\_\_.

7. This law states that \_\_\_\_\_.

**Write the mole ratio in the space provided and solve each problem.**



8. How many moles of sodium hydroxide are produced when 1.50 mol sodium peroxide reacts with water?

$$1.50 \text{ mol Na}_2\text{O}_2 \quad \times \quad \frac{4 \text{ mol NaOH}}{2 \text{ mol Na}_2\text{O}_2} \quad = \quad \underline{3.00} \quad \text{mol NaOH}$$

9. How many moles of water are required to produce 2.15 mol oxygen gas in this reaction?

$$2.15 \text{ mol O}_2 \quad \times \quad \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} \quad = \quad \underline{4.30} \quad \text{mol H}_2\text{O}$$

**Stoichiometry Practice Problems #1..... Solve each problem, show all work.**

1. How many **grams** of  $\text{Cl}_2$  must be reacted with excess  $\text{NaI}$  if 10.0g of  $\text{NaCl}$  are needed?



# Stoichiometry Homework Answer Key Chemfiesta

**L Towner**



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