

# Stoichiometry Limiting Reagent

## Answers

- 1 Consider the following reaction:



- (a) How many grams of  $\text{NH}_3$  can be produced from the reaction of 28 g of  $\text{N}_2$  and 25 g of  $\text{H}_2$ ?

$$28 \text{ g N}_2 \times \frac{1 \text{ mol N}_2}{28 \text{ g N}_2} \times \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} \times \frac{17 \text{ g NH}_3}{1 \text{ mol NH}_3} = 34 \text{ g NH}_3$$

$$25 \text{ g H}_2 \times \frac{1 \text{ mol H}_2}{2 \text{ g H}_2} \times \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \times \frac{17 \text{ g NH}_3}{1 \text{ mol NH}_3} = 142 \text{ g NH}_3$$

34 g of  $\text{NH}_3$  can be produced.

- (b) How much of the excess reagent is left over?

$$28 \text{ g N}_2 \times \frac{1 \text{ mol N}_2}{28 \text{ g N}_2} \times \frac{3 \text{ mol H}_2}{1 \text{ mol N}_2} \times \frac{2 \text{ g H}_2}{1 \text{ mol H}_2} = 6 \text{ g H}_2 \text{ used}$$

$$25 \text{ g H}_2 - 6 \text{ g H}_2 = 19 \text{ g H}_2 \text{ remains}$$

- 2 Consider the following reaction:



What volume of  $\text{H}_2$  at STP is produced from the reaction of 50 g of  $\text{Mg}$  and 75 g of  $\text{HCl}$ ?

$$50 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.3 \text{ g Mg}} \times \frac{1 \text{ mol H}_2}{1 \text{ mol Mg}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} = 46 \text{ L H}_2$$

$$75 \text{ g HCl} \times \frac{1 \text{ mol HCl}}{36.5 \text{ g HCl}} \times \frac{1 \text{ mol H}_2}{2 \text{ mol HCl}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} = 23 \text{ L H}_2$$

23 L  $\text{H}_2$  is produced.

- 3 Zinc and sulfur react to form zinc sulfide according to the following equation:



Suppose 25 g of zinc and 30 g of sulfur react. Which reagent is limiting?

$$25 \text{ g Zn} \times \frac{1 \text{ mol Zn}}{65.4 \text{ g Zn}} \times \frac{1 \text{ mol ZnS}}{1 \text{ mol Zn}} \times \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 37.3 \text{ g ZnS}$$

$$30 \text{ g S} \times \frac{1 \text{ mol S}}{32 \text{ g S}} \times \frac{1 \text{ mol ZnS}}{1 \text{ mol S}} \times \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 91.2 \text{ g ZnS}$$

Zn is the limiting reagent.

# Stoichiometry Limiting Reagent Pg 66 Answers

**Julie C. Gilbert**



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### **Table of Contents Stoichiometry Limiting Reagent Pg 66 Answers**

1. Understanding the eBook Stoichiometry Limiting Reagent Pg 66 Answers
  - The Rise of Digital Reading Stoichiometry Limiting Reagent Pg 66 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Limiting Reagent Pg 66 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Stoichiometry Limiting Reagent Pg 66 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Limiting Reagent Pg 66 Answers
  - Personalized Recommendations

- Stoichiometry Limiting Reagent Pg 66 Answers User Reviews and Ratings
- Stoichiometry Limiting Reagent Pg 66 Answers and Bestseller Lists
- 5. Accessing Stoichiometry Limiting Reagent Pg 66 Answers Free and Paid eBooks
  - Stoichiometry Limiting Reagent Pg 66 Answers Public Domain eBooks
  - Stoichiometry Limiting Reagent Pg 66 Answers eBook Subscription Services
  - Stoichiometry Limiting Reagent Pg 66 Answers Budget-Friendly Options
- 6. Navigating Stoichiometry Limiting Reagent Pg 66 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Stoichiometry Limiting Reagent Pg 66 Answers Compatibility with Devices
  - Stoichiometry Limiting Reagent Pg 66 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Stoichiometry Limiting Reagent Pg 66 Answers
  - Highlighting and Note-Taking Stoichiometry Limiting Reagent Pg 66 Answers
  - Interactive Elements Stoichiometry Limiting Reagent Pg 66 Answers
- 8. Staying Engaged with Stoichiometry Limiting Reagent Pg 66 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Stoichiometry Limiting Reagent Pg 66 Answers
- 9. Balancing eBooks and Physical Books Stoichiometry Limiting Reagent Pg 66 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Stoichiometry Limiting Reagent Pg 66 Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Stoichiometry Limiting Reagent Pg 66 Answers
  - Setting Reading Goals Stoichiometry Limiting Reagent Pg 66 Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Stoichiometry Limiting Reagent Pg 66 Answers
  - Fact-Checking eBook Content of Stoichiometry Limiting Reagent Pg 66 Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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