

Stoichiometry Practice 2.1 - Answer Key

Objectives: - Use stoichiometry to calculate the mass of any substance involved in a chemical reaction

Directions: Answer the following stoichiometry problems and show your work.

1. If a piece of magnesium with a mass of 2.76 g is added to a solution of hydrochloric acid (HCl), what mass of H_2 would be produced? **[0.229 g]**



$$2.76 \text{ g Mg} \times \left(\frac{1 \text{ mol Mg}}{24.305 \text{ g Mg}} \right) \left(\frac{1 \text{ mol } H_2}{1 \text{ mol Mg}} \right) \left(\frac{2.02 \text{ g } H_2}{1 \text{ mol } H_2} \right) = \boxed{0.229 \text{ g } H_2}$$

2. What mass of barium chloride is needed to react completely with 46.8 g of sodium phosphate?



$$46.8 \text{ g } Na_3PO_4 \times \left(\frac{1 \text{ mol } Na_3PO_4}{163.94 \text{ g } Na_3PO_4} \right) \left(\frac{3 \text{ mol } BaCl_2}{2 \text{ mol } Na_3PO_4} \right) \left(\frac{208.23 \text{ g } BaCl_2}{1 \text{ mol } BaCl_2} \right) = \boxed{89.1 \text{ g } BaCl_2}$$

3. In the reaction, where magnesium oxide reacts with water to produce magnesium hydroxide. What mass of magnesium oxide is needed to produce 264 g of magnesium hydroxide?



$$264 \text{ g } Mg(OH)_2 \times \left(\frac{1 \text{ mol } Mg(OH)_2}{58.32 \text{ g } Mg(OH)_2} \right) \left(\frac{1 \text{ mol } MgO}{1 \text{ mol } Mg(OH)_2} \right) \left(\frac{40.30 \text{ g } MgO}{1 \text{ mol } MgO} \right) = \boxed{18.2 \text{ g } MgO}$$

4. Methane (CH_4) burns in oxygen to produce carbon dioxide and water. What mass of oxygen would be necessary for the reaction if you had 15.0 g of methane?



$$15.0 \text{ g } CH_4 \times \left(\frac{1 \text{ mol } CH_4}{16.04 \text{ g } CH_4} \right) \left(\frac{2 \text{ mol } O_2}{1 \text{ mol } CH_4} \right) \left(\frac{32.00 \text{ g } O_2}{1 \text{ mol } O_2} \right) = \boxed{59.8 \text{ g } O_2}$$

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Table of Contents **Stoichiometry Mass Mass With Answers**

1. Understanding the eBook **Stoichiometry Mass Mass With Answers**
 - The Rise of Digital Reading **Stoichiometry Mass Mass With Answers**
 - Advantages of eBooks Over Traditional Books
2. Identifying **Stoichiometry Mass Mass With Answers**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **Stoichiometry Mass Mass With Answers**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **Stoichiometry Mass Mass With Answers**
 - Personalized Recommendations
 - **Stoichiometry Mass Mass With Answers** User Reviews and Ratings
 - **Stoichiometry Mass Mass With Answers** and Bestseller Lists

5. Accessing Stoichiometry Mass Mass With Answers Free and Paid eBooks
 - Stoichiometry Mass Mass With Answers Public Domain eBooks
 - Stoichiometry Mass Mass With Answers eBook Subscription Services
 - Stoichiometry Mass Mass With Answers Budget-Friendly Options
6. Navigating Stoichiometry Mass Mass With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Mass Mass With Answers Compatibility with Devices
 - Stoichiometry Mass Mass With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Mass Mass With Answers
 - Highlighting and Note-Taking Stoichiometry Mass Mass With Answers
 - Interactive Elements Stoichiometry Mass Mass With Answers
8. Staying Engaged with Stoichiometry Mass Mass With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Mass Mass With Answers
9. Balancing eBooks and Physical Books Stoichiometry Mass Mass With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Mass Mass With Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Mass Mass With Answers
 - Setting Reading Goals Stoichiometry Mass Mass With Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry Mass Mass With Answers
 - Fact-Checking eBook Content of Stoichiometry Mass Mass With Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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