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CLASS:

INTRODUCTION

These problems involve stoichiometric calculations where one or more of the reactants or products are in solution.

No.	Question	Answer
1	Determine the number of mol of barium sulfate that will be precipitated when 200.0 mL of 0.450 M barium nitrate solution is added to an excess of sodium sulfate solution, given the equation for the reaction is: Ba(NO ₃) ₂ (aq) + Na ₂ SO ₂ (aq) → BaSO ₂ (s) + 2NaNO ₃ (aq)	
2	Liquid hydrazine reacts explosively with hydrogen peroxide according to the equation: N ₂ H ₂ (1) + 2H ₂ O ₂ (aq) → N ₂ (g) + 4H ₂ O(g) Determine the number of mol of gaseous products that would be formed by the reaction of 3.50 L of 4.30 M hydrogen peroxide with excess hydrazine.	
3	17.0 mL of 0.880 M lead(II) nitrate is added to excess sodium hydroxide solution to produce a precipitate of lead(II) hydroxide. Calculate the mass of precipitate formed, given that the equation for this reaction is: Pb(NO ₃) ₂ (aq) + 2NaOH(aq) → Pb(OH) ₂ (s) + 2NaNO ₃ (aq)	

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