

# Stoichiometry Worksheet and Key

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1. How many moles of  $\text{O}_2$  will be formed from 1.65 moles of  $\text{KClO}_3$ ?

1.65 mol $\text{KClO}_3$	_____ mol $\text{O}_2$	= _____ mol $\text{O}_2$
	_____ mol $\text{KClO}_3$	

2. How many moles of  $\text{KClO}_3$  are needed to make 3.50 moles of  $\text{KCl}$ ?

3.50 mol $\text{KCl}$		= _____ mol $\text{KClO}_3$

3. How many moles of  $\text{KCl}$  will be formed from 2.73 moles of  $\text{KClO}_3$ ?

_____		= _____



4. How many moles of  $\text{Fe}_2\text{O}_3$  are produced when 0.275 moles of  $\text{Fe}$  is reacted?

0.275 mol $\text{Fe}$		= _____ mol $\text{Fe}_2\text{O}_3$

5. How many moles of  $\text{Fe}_2\text{O}_3$  are produced when 31.0 moles of  $\text{O}_2$  is reacted?

_____		= _____

6. How many moles of  $\text{O}_2$  are needed to react with 8.9 moles of  $\text{Fe}$ ?

_____		= _____



7. How many moles of  $\text{O}_2$  are produced when 1.26 moles of  $\text{H}_2\text{O}$  is reacted?

8. How many moles of  $\text{H}_2\text{O}$  are needed to produce 55.7 moles of  $\text{H}_2$ ?

9. If enough  $\text{H}_2\text{O}$  is reacted to produce 3.40 moles of  $\text{H}_2$ , then how many moles of  $\text{O}_2$  must have been made?  
(a bit challenging, but just think about it and you can probably figure it out)

# Stoichiometry Problem Sheet 1

**George W. Jr. Latimer**



## **Stoichiometry Problem Sheet 1:**

**A Visual Analogy Guide to Chemistry, 2e** Paul A Krieger, 2018-02-01 A Visual Analogy Guide to Chemistry is the latest in the innovative and widely used series of books by Paul Krieger This study guide delivers a big picture view of difficult concepts and effective study tools to help students learn and understand the details of general organic and biochemistry topics A Visual Analogy Guide to Chemistry is a worthwhile investment for any introductory chemistry student **Jacaranda Chemistry 1 VCE Units 1 and 2, learnON and Print** Neale Taylor, Angela Stubbs, Robert Stokes, 2022-11-30 Developed by expert Victorian teachers for VCE students The NEW Jacaranda Chemistry VCE series continues to deliver curriculum aligned material that caters to students of all abilities Our expert author team of practising teachers and assessors ensures 100% coverage of the new VCE Chemistry Study Design 2023 2027 **Elements of Environmental Chemistry** Ronald A. Hites, Jonathan D. Raff, 2013-08-26 From Reviews of the First Edition This splendid at times humorous and reasonably priced little book has much to commend it to undergraduate chemists and to other science students J G Farmer University of Edinburgh Complex environmental issues are presented in simple terms to help readers grasp the basics and solve relevant problems J Albaiges University of Barcelona The main strength of the book lies in its explanations of the calculation of quantitative relationships Each chapter includes 15 20 problems that are carefully chosen from a didactic standpoint for which the reader can find solutions at the end D Lenoir Institute for Ecological Chemistry What drew me to the first edition was the style the no nonsense down to earth explanations and the practical examples that litter the text The dry humor expressed in the footnotes is great and reminds me of other classic texts T Clough Lincoln University A practical approach to environmental chemistry Providing readers with the fundamentals of environmental chemistry and a toolbox for putting them into practice Elements of Environmental Chemistry Second Edition is a concise accessible and hands on volume designed for students and professionals working in the chemical and environmental sciences Tutorial in style this book fully incorporates real world problems and extensive end of chapter problem sets to immerse the reader in the field Chapters cover mass balance chemical kinetics carbon dioxide equilibria pesticide structures and much more Extensively revised updated and expanded this Second Edition includes new chapters on atmospheric chemistry climate change and polychlorinated biphenyls and dioxins and brominated flame retardants In addition new practice problems and a helpful tutorial on organic chemistry names and structures have been added to improve both the scope and accessibility of the book **STOICHIOMETRY** NARAYAN CHANGDER, 2024-04-01 Note Anyone can request the PDF version of this practice set workbook by emailing me at cbsenet4u gmail com You can also get full PDF books in quiz format on our youtube channel <https://www.youtube.com/smartquiziz> I will send you a PDF version of this workbook This book has been designed for candidates preparing for various competitive examinations It contains many objective questions specifically designed for different exams Answer keys are provided at the end of each page It will undoubtedly serve as the best preparation material for aspirants This book is an

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*Resources in Education* ,1994     *Food Protein Chemistry* Joe Regenstein,Carrie Regenstein,1984-08-28 Food Protein Chemistry An Introduction for Food Scientists discusses food proteins and how they are studied Proteins are both biological entities and physicochemical compounds and they will be examined in both contexts in this volume The chemical and physical properties of proteins will be viewed from the perspective of chemists despite the fact that their use in the food supply emphasizes their biological nature Key topics discussed include proteins as essential to life amino acids protein classification selected proteins of the most important food systems and protein structure The book also includes chapters on protein measurement protein purification and spectral techniques for the study of proteins The book requires readers to have the equivalent of the Institute of Food Technologists requirements for undergraduate food science majors It also assumes a knowledge of math through calculus While primarily intended for senior and first year graduate food science students the text may also be useful to researchers in allied fields     Chemistry, Student Study Guide James E. Brady,Fred Senese,2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH<sub>4</sub> from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it

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