Stoichiometry Worksheet and Key

2 KClO₃ → 2 KCl + 3 O₂

How many moles of O₂ will be formed from 1.65 moles of KClO₃?

 $\begin{array}{c|cccc}
1.65 & \text{mol KClO}_3 & \underline{\qquad} & \text{mol O}_2 \\
& \underline{\qquad} & \text{mol KClO}_3 & \underline{\qquad} & \text{mol O}_2
\end{array}$

2. How many moles of KClO₃ are needed to make 3.50 moles of KCl?

3.50mol KCl = ___ mol KClO₃

3. How many moles of KCl will be formed from 2.73 moles of KClO₃?

4 Fe + 3 O₂ → 2 Fe₂O₃

4. How many moles of Fe₂O₃ are produced when 0.275 moles of Fe is reacted?

0.275 mol Fe = ___ mol Fe₂O₃

5. How many moles of Fe₂O₃ are produced when 31.0 moles of O₂ is reacted?

6. How many moles of O₂ are needed to react with 8.9 moles of Fe?

 $2 H_2O \rightarrow 2 H_2 + O_2$

7. How many moles of O₂ are produced when 1.26 moles of H₂O is reacted?

How many moles of H₂O are needed to produce 55.7 moles of H₂?

If enough H₂O is reacted to produce 3.40 moles of H₂, then how may moles of O₂ must have been made?
 (a bit challenging, but just think about it and you can probably figure it out)

Stoichiometry Problem Solving Workbook Answers

Loretta Jones, Carl Hoeger, Peter William Atkins, Regina Schoenfield-Tacher

Stoichiometry Problem Solving Workbook Answers:

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