

# CHAPTER 9 REVIEW

## Stoichiometry

KEY

### SECTION 3

**PROBLEMS** Write the answer on the line to the left. Show all your work in the space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

$$\% \text{ yield} = \frac{22 \text{ g}}{25 \text{ g}} \times 100 = 88\%$$

2. 6.0 mol of  $\text{N}_2$  are mixed with 12.0 mol of  $\text{H}_2$  according to the following equation:



2.0 mol x 3  $\text{N}_2$

ratio



6.0 mol  $\text{N}_2$

12.0 mol  $\text{H}_2$

- a. Which chemical is in excess? What is the excess in moles?

$$\text{N}_2: \left( \frac{2 \text{ mol}}{1 \text{ H}_2} \right) \left( \frac{1 \text{ N}_2}{3 \text{ H}_2} \right) = 9.0 \text{ mol}$$

not  $3 \times \text{N}_2 \Rightarrow \text{H}_2$  is limiting

8.0 mol  $\text{NH}_3$

- b. Theoretically, how many moles of  $\text{NH}_3$  will be produced?

$$\text{mol NH}_3 = \left( \frac{12.0 \text{ mol}}{3 \text{ H}_2} \right) \left( \frac{2 \text{ NH}_3}{1 \text{ N}_2} \right) = 8.0 \text{ mol}$$

6.4 mol  $\text{NH}_3$

- c. If the percentage yield of  $\text{NH}_3$  is 80%, how many moles of  $\text{NH}_3$  are actually produced?

$$\% \text{ yield} = \frac{\text{actual}}{\text{theor}} \times 100$$

$$\text{actual} = \frac{(\% \text{ yield})(\text{theor})}{100} = \frac{(80)(8.0)}{100} = 6.4 \text{ mol}$$

3. 0.050 mol of  $\text{Ca}(\text{OH})_2$  are combined with 0.080 mol of  $\text{HCl}$  according to the following equation:



0.10 mol  $\text{HCl}$

- a. How many moles of  $\text{HCl}$  are required to neutralize all 0.050 mol of  $\text{Ca}(\text{OH})_2$ ?

$$\text{mol HCl} = \left( \frac{0.050 \text{ mol}}{1 \text{ Ca}(\text{OH})_2} \right) \left( \frac{2 \text{ HCl}}{1 \text{ Ca}(\text{OH})_2} \right) = 0.10 \text{ mol}$$

limiting

# Stoichiometry Section 9 Answers

**John T. Moore, Chris Hren, Peter J.  
Mikulecky**



## **Stoichiometry Section 9 Answers:**

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## **Table of Contents Stoichiometry Section 9 Answers**

1. Understanding the eBook Stoichiometry Section 9 Answers
  - The Rise of Digital Reading Stoichiometry Section 9 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Section 9 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Stoichiometry Section 9 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Section 9 Answers
  - Personalized Recommendations
  - Stoichiometry Section 9 Answers User Reviews and Ratings
  - Stoichiometry Section 9 Answers and Bestseller Lists

5. Accessing Stoichiometry Section 9 Answers Free and Paid eBooks
  - Stoichiometry Section 9 Answers Public Domain eBooks
  - Stoichiometry Section 9 Answers eBook Subscription Services
  - Stoichiometry Section 9 Answers Budget-Friendly Options
6. Navigating Stoichiometry Section 9 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Stoichiometry Section 9 Answers Compatibility with Devices
  - Stoichiometry Section 9 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Stoichiometry Section 9 Answers
  - Highlighting and Note-Taking Stoichiometry Section 9 Answers
  - Interactive Elements Stoichiometry Section 9 Answers
8. Staying Engaged with Stoichiometry Section 9 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Stoichiometry Section 9 Answers
9. Balancing eBooks and Physical Books Stoichiometry Section 9 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Stoichiometry Section 9 Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Section 9 Answers
  - Setting Reading Goals Stoichiometry Section 9 Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry Section 9 Answers
  - Fact-Checking eBook Content of Stoichiometry Section 9 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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