

Stoichiometry Study Guide

Below are questions that will help you study for your test. Not everything on your test is represented in the study guide, and not all of the study guide questions will be on your test. (Please note that I will not give you the formula for **ANYTHING** on your test. Everything will be in terms of the name of a compound. I.e., you will be given sodium hydroxide **NOT** NaOH.) All answers should be recorded on a separate sheet of paper!

1. Compare and contrast weight and mass.
2. Define molecular mass. What are the units of molecular mass?
3. Calculate the molecular mass of each of the following.
 - a. carbon dioxide
 - b. carbon tetrachloride
 - c. magnesium oxide
 - d. copper (II) chloride
 - e. hydroxide
 - f. sodium
 - g. hydrogen sulfide
 - h. potassium permanganate
4. Given the following information, write a chemical equation, balance it, and answer the stoichiometry questions. calcium fluoride reacts with rubidium oxide to produce calcium oxide and rubidium fluoride.
 - a. How many moles of calcium fluoride are required to produce 56 moles of rubidium fluoride?
 - b. How many grams of rubidium fluoride will be produced from 4 moles of rubidium oxide?
 - c. How many grams of calcium fluoride would be required to produce 45.23 grams of rubidium fluoride?
 - d. How many grams of rubidium oxide are required to produce 12.24 moles of rubidium fluoride?
 - e. How many moles of calcium oxide will be produced from 1.25 moles of calcium fluoride?
5. Compare and contrast limiting and excess reagents.
6. Given the following information, write a chemical equation, balance it, and answer the stoichiometry and limiting reagent questions. strontium reacts with oxygen to produce strontium oxide.
 - a. What is the limiting reagent in the equation above if 14.75 g of each reactant is used? (be sure to show all of your work)
 - b. What is the maximum amount of strontium oxide that can be produced if 14.75 g of each reactant is used? (be sure to show all of your work)
 - c. How many grams of excess reagent will be left once the reaction is complete?
7. Given the following information, write a chemical equation, balance it, and answer the stoichiometry and limiting reagent questions. sodium hydroxide reacts with hydrochloric acid to produce sodium chloride and water.
 - a. What is the limiting reagent in the equation above if 14.75 g of each reactant is used? (be sure to show all of your work)
 - b. What is the maximum amount of strontium oxide that can be produced if 14.75 g of each reactant is used? (be sure to show all of your work)
 - c. How many grams of excess reagent will be left once the reaction is complete?

Stoichiometry Study Guide Answers Chemistry

Martha Mackin



Stoichiometry Study Guide Answers Chemistry:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes. Also includes solutions and explanations to odd-numbered exercises.

Chemistry, Student Study Guide

James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane (CH_4) from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement.

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18-chapter text designed to be used with Basics for Chemistry textbook. Each chapter contains Overview, Topical Outline, Skills, and Common Mistakes, which are all keyed to the textbook for easy cross-reference. The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms, a summary of general concepts, and a list of numerical exercises, while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook. The Fill-in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter, and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples. The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry, while the Practical Test section includes matching and multiple

choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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