

Answers to Exercise 1.4 Stoichiometry Basics

Tables like the ones in this answer key are enormously helpful at keeping you organized! Sometimes, they even suggest the next logical step in a calculation...

1.	$C_3H_8(g)$	+	$5 O_2(g)$	$\rightarrow$	$4 H_2O(g)$	+	$3 CO_2(g)$
M	44.096 g/mol		31.9988 g/mol		18.0152 g/mol		44.010 g/mol
$m_{initial}$	25 g						
$n_{initial}$	0.57 mol		2.8 mol		0 mol		0 mol
n_{change}	-0.57 mol		-2.8 mol		+2.3 mol		+1.7 mol
n_{final}	0 mol		0 mol		2.3 mol		1.7 mol
$m_{initial}$	25 g		91 g		0 g		0 g
m_{final}	0 g		0 g		41 g		75 g

Step 1: Write a balanced chemical equation for the reaction

see above

Step 2: Organize all known information

see above; this involves reporting masses, concentrations, pressures, etc. as well as calculating molar masses, checking the data sheet for relevant densities, etc.

Step 3: Calculate moles of C_3H_8 reacted ($n_{initial}$)

$$n_{C_3H_8-initial} = 25 \text{ g } C_3H_8 \times \frac{1 \text{ mol}}{44.096 \text{ g}} = 0.57 \text{ mol } C_3H_8$$

Step 4: Use mole ratio to calculate moles of O_2 required for reaction ($n_{initial}$)

$$n_{O_2-initial} = 0.57 \text{ mol } C_3H_8 \times \frac{5 \text{ mol } O_2}{1 \text{ mol } C_3H_8} = 2.8 \text{ mol } O_2$$

Step 5: Calculate mass of O_2 required for reaction ($m_{initial}$)

$$m_{O_2-initial} = 2.8 \text{ mol } O_2 \times \frac{31.9988 \text{ g}}{1 \text{ mol}} = 91 \text{ g } O_2 \quad \text{answer to (a)}$$

Step 6: Use mole ratio to calculate moles of H_2O and CO_2 produced (n_{final})

$$n_{H_2O-final} = 0.57 \text{ mol } C_3H_8 \times \frac{4 \text{ mol } H_2O}{1 \text{ mol } C_3H_8} = 2.3 \text{ mol } H_2O$$

$$n_{CO_2-final} = 0.57 \text{ mol } C_3H_8 \times \frac{3 \text{ mol } CO_2}{1 \text{ mol } C_3H_8} = 1.7 \text{ mol } CO_2$$

Step 7: Calculate masses of H_2O and CO_2 produced (m_{final})

$$m_{H_2O-final} = 2.3 \text{ mol } H_2O \times \frac{18.0152 \text{ g}}{1 \text{ mol}} = 41 \text{ g } H_2O \quad \text{answer to (b)}$$

$$m_{CO_2-final} = 1.7 \text{ mol } CO_2 \times \frac{44.010 \text{ g}}{1 \text{ mol}} = 75 \text{ g } CO_2 \quad \text{answer to (c)}$$

Step 8: Check your work

Does your answer seem reasonable? Are sig. fig. correct?

25 g + 91 g = 41 g + 75 g = 116 g so the Law of Conservation of Mass is obeyed. Therefore, this seems like a reasonable set of answers.

answer to (d)

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Table of Contents Stoichiometry Test Questions And Answers

1. Understanding the eBook Stoichiometry Test Questions And Answers
 - The Rise of Digital Reading Stoichiometry Test Questions And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Test Questions And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Test Questions And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Test Questions And Answers
 - Personalized Recommendations
 - Stoichiometry Test Questions And Answers User Reviews and Ratings

- Stoichiometry Test Questions And Answers and Bestseller Lists
- 5. Accessing Stoichiometry Test Questions And Answers Free and Paid eBooks
 - Stoichiometry Test Questions And Answers Public Domain eBooks
 - Stoichiometry Test Questions And Answers eBook Subscription Services
 - Stoichiometry Test Questions And Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Stoichiometry Test Questions And Answers Compatibility with Devices
 - Stoichiometry Test Questions And Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Test Questions And Answers
 - Highlighting and Note-Taking Stoichiometry Test Questions And Answers
 - Interactive Elements Stoichiometry Test Questions And Answers
- 8. Staying Engaged with Stoichiometry Test Questions And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Test Questions And Answers
- 9. Balancing eBooks and Physical Books Stoichiometry Test Questions And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Test Questions And Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Stoichiometry Test Questions And Answers
 - Setting Reading Goals Stoichiometry Test Questions And Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Stoichiometry Test Questions And Answers
 - Fact-Checking eBook Content of Stoichiometry Test Questions And Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web junie lewensoriëntering eksamen vir graad 9 leerders dis opgestel volgens die spesifikasies van die departement van onderwys die eksamen sluit n verskeidenheid

wiskunde graad 9 tweede kwartaal vraestel en memo litnet - Mar 06 2023

web g09 wiskunde v1 wiskunde junie eksamen 2021 vraestel 1 graad 9 eksaminator s kotzé tyd 1 uur moderator k britz totaal 50

vraag 1 litnet - Sep 12 2023

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2 4 12 vraag 7 7 1 los op vir □ 4 7 2

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web hierdie produk bevat die volgende graad 12 vraestelle met memorandums vir die junie eksamens let wel vraestelle is dieselfde as die individuele vraestelle beskikbaar moet