




Stoichiometric Calculations

Equation:	$2 \text{H}_2(\text{g})$	+	$\text{O}_2(\text{g})$	\longrightarrow	$2 \text{H}_2\text{O}(\text{l})$
Molecules:	2 molecules H_2	+	1 molecule O_2	\longrightarrow	2 molecules H_2O
					
Mass (amu):	4.0 amu H_2	+	32.0 amu O_2	\longrightarrow	36.0 amu H_2O
Amount (mol):	2 mol H_2	+	1 mol O_2	\longrightarrow	2 mol H_2O
Mass (g):	4.0 g H_2	+	32.0 g O_2	\longrightarrow	36.0 g H_2O

The coefficients in the balanced equation give the ratio of *moles* of reactants and products

Stoichiometric Calculations Study Guide

James E. Brady, Fred Senese



Stoichiometric Calculations Study Guide:

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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