

## Chapter 9 Stoichiometry: Practice Test

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*Solve the following problems in the space provided. Show your work.*

- 1) a. What is the limiting reagent when 50.2 g of nitrogen react with 10.7 g of hydrogen according to this balanced equation?  
b. What mass of ammonia would be produced?



- 2) How many grams of CO are needed to react with an excess of  $\text{Fe}_2\text{O}_3$  to produce 558 g of Fe? The unbalanced equation for the equation is:



- 3) How many grams of butane ( $\text{C}_4\text{H}_{10}$ ) must be burned in an excess of  $\text{O}_2$  to produce 15.0 g of  $\text{CO}_2$ ?



- 4) a. If 4.0 g of  $\text{H}_2$  are made to react with excess CO, how many grams of  $\text{CH}_3\text{OH}$  can theoretically be produced according to the following equation?



- b. If 28.0 g of  $\text{CH}_3\text{OH}$  are actually produced, what is the percent yield?

# Stoichiometry 12 Test A Answers

**Kaplan Test Prep**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the Kaplan Test Prep text.

## **Stoichiometry 12 Test A Answers:**

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