

Stoichiometry Review: Test Review



1. What is stoichiometry? [] Study of quantitative relationships in chemical reaction.
2. On which law is stoichiometry based? [] Law of conservation of mass.
3. According to the law of conservation of mass, the mass of the [] ^{reactants} must [equal] the mass of the ^{products}.
4. What is a mole ratio? [] fractions/ratio comparing two substances.
5. Given the following balanced equation: $2\text{Al} + 3\text{Br}_2 \rightarrow 2\text{AlBr}_3$, determine the mole ratio used for calculating moles of aluminum bromide from moles of bromine. [] $2\text{mol AlBr}_3 / 3\text{mol Br}_2$
6. What must you have before beginning any stoichiometric problem? [] balanced equation
7. What is a limiting reactant? [] runs out
8. What is an excess reactant? [] left over
9. A booklet is prepared from 2 covers, 3 staples, and 20 pages. Assume we have 60 covers, 120 staples, and 400 pages.
 - a. What is the largest number of complete booklets that can be made? [20]
 - b. Which is the "limiting reactant" in this system? [pages]
10. How many moles of lithium chloride will be formed by the reaction of chlorine with 0.046 mol of lithium bromide in the following reaction?
 $[2] \text{LiBr} + [1] \text{Cl}_2 \rightarrow [2] \text{LiCl} + [1] \text{Br}_2$

$$\frac{0.046 \text{ mol LiBr}}{2 \text{ mol LiBr}} \times \frac{2 \text{ mol LiCl}}{2 \text{ mol LiBr}} = 0.046 \text{ mol LiCl}$$

11. Hydrogen peroxide breaks down, releasing oxygen, in the following reaction. What mass of oxygen is produced when 1.840 mol of H_2O_2 decomposes?

$$\frac{1.840 \text{ mol H}_2\text{O}_2}{2 \text{ mol H}_2\text{O}_2} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}_2} \times \frac{32.09 \text{ g O}_2}{1 \text{ mol O}_2} = 29.44 \text{ g O}_2$$

12. Hydrogen is generated by passing hot steam over iron, which oxidizes to form Fe_2O_3 , in the following equation. If 625 g of Fe_2O_3 is produced in the reaction, how many moles of hydrogen are produced at the same time?

Stoichiometry Form A Test Answer Key

Peter Atkins



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Prevention and Control Paul Mac Berthouex, Linfield C. Brown, 2017-10-04 This book examines how chemistry chemical processes and transformations are used for pollution prevention and control Pollution prevention reduces or eliminates pollution at the source whereas pollution control involves destroying reducing or managing pollutants that cannot be

eliminated at the source Applications of environmental chemistry are further illustrated by nearly 150 figures numerous example calculations and several case studies designed to develop analytical and problem solving skills The book presents a variety of practical applications and is unique in its integration of pollution prevention and control as well as air water and solid waste management **Problems Of Instrumental Analytical Chemistry: A Hands-on Guide (Second Edition)** Jose Manuel Andrade Garda,Alatzne Carlosena-zubieta,Marcos Gestal-pose,Maria Paz Gomez-carracedo,Miguel A Maestro-saavedra,Maria Del Carmen Prieto-blanco,Rosa Maria Soto-ferreiro,Javier Teran-baamonde,2025-07-02 The complex field of analytical chemistry requires knowledge and application of the fundamental principles of numerical calculation Problems of Instrumental Analytical Chemistry provides support and guidance to help students develop these numerical strategies to generate information from experimental results in an efficient and reliable way The book contains exercises that provide standard protocols for the most common calculations in the daily work of a laboratory Also included are easy to follow diagrams to facilitate understanding and avoid common errors making this textbook perfect as a hands on accompaniment to in class learning The subjects covered follow a course in analytical chemistry from the initial basics of data analysis to applications of mass UV VIS infrared and atomic spectrometry and chromatography concluding with an overview of nuclear magnetic resonance and electrochemistry Intended as a self training tool for undergraduates in chemistry analytical chemistry and related subjects this book is also useful as a reference for scientists looking to brush up on their knowledge of instrumental techniques in laboratories This second edition builds upon the first with new and updated content as well as QR codes distributed throughout directing readers to dedicated materials and websites hosting additional information examples and models **Tests** Richard C. Sweetland,1991 Provides descriptive information on some 3 200 tests for the benefit of test evaluators and others who need to determine if a test suits their purposes Each entry includes test name and author an indication of the population for which the test is intended how the test is administered major features of the test how the test is timed and scored pricing and availability information and publisher and distributor Cross referenced and indexed eight ways Distributed by Gale Research Annotation copyrighted by Book News Inc Portland OR **Journal of Research of the National Bureau of Standards** United States. National Bureau of Standards,1953-07

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